

WEAK ACIDS, BASES, AND THEIR SALTS

INTRODUCTION

The purpose of this experiment is to determine the pH of various concentrations of acids, bases, and salts. The data will be used to determine the acid–base dissociation constants of the substances being investigated. The variation of the degree of ionization of a weak acid and of a weak base with concentration also will be determined.

STRONG ACIDS AND STRONG BASES

When a strong acid such as HCl or a strong base such as NaOH is dissolved in water, it dissociates completely into ions as shown in the following two equations:



As a result, the concentration of hydronium ions, $[\text{H}_3\text{O}^+]$, or the concentration of hydroxide ions, $[\text{OH}^-]$, essentially is equal to the concentration of the strong acid or strong base, respectively. Hydronium ion concentrations commonly range from 1.0 M to 10^{-14} M and frequently are expressed on the logarithmic scale of pH.

Eq. 3 $\text{pH} = -\log[\text{H}_3\text{O}^+]$

The $[\text{OH}^-]$ is related to $[\text{H}_3\text{O}^+]$ by the ion product of water, K_w , which at 25 °C has the value

Eq. 4 $K_w = [\text{OH}^-][\text{H}_3\text{O}^+] = 1.0 \times 10^{-14}$

The relationship between pH and pOH ($-\log[\text{OH}^-]$) is obtained from Eq. 4.

$$\text{pOH} = -\log[\text{OH}^-]$$

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