

Chapter 12: Chemical Equilibrium Quiz

1. The equilibrium between carbon dioxide gas and carbonic acid is very important in biology and environmental science.



Which one of the following is the correct equilibrium constant expression (K_c) for this reaction?

- A) $K_c = \frac{[\text{H}_2\text{CO}_3]}{[\text{CO}_2][\text{H}_2\text{O}]}$ D) $K_c = \frac{[\text{CO}_2]}{[\text{H}_2\text{CO}_3]}$
 B) $K_c = \frac{[\text{CO}_2][\text{H}_2\text{O}]}{[\text{H}_2\text{CO}_3]}$ E) $K_c = \frac{1}{[\text{H}_2\text{CO}_3]}$
 C) $K_c = \frac{[\text{H}_2\text{CO}_3]}{[\text{CO}_2]}$

2. The following reactions occur at 500 K. Arrange them in order of increasing tendency to proceed to completion (least \rightarrow greatest tendency).



- A) $2 < 1 < 3$ D) $3 < 2 < 1$
 B) $1 < 2 < 3$ E) $3 < 1 < 2$
 C) $2 < 3 < 1$

3. Consider the two gaseous equilibria:



The values of the equilibrium constants K_1 and K_2 are related by

- A) $K_2 = (K_1)^2$ D) $K_2 = 1/K_1$
 B) $(K_2)^2 = K_1$ E) none of these.
 C) $K_2 = 1/(K_1)^2$

4. When the following reaction is at equilibrium, which of these relationships is *always* true?



- A) $[\text{NO}][\text{Cl}_2] = [\text{NOCl}]$ D) $2[\text{NO}] = [\text{Cl}_2]$
 B) $[\text{NO}]^2[\text{Cl}_2] = [\text{NOCl}]^2$ E) $[\text{NO}]^2[\text{Cl}_2] = K_c[\text{NOCl}]^2$
 C) $[\text{NOCl}] = [\text{NO}]$

5. Calculate K_c for the reaction $2\text{HI}(\text{g}) \rightleftharpoons \text{H}_2(\text{g}) + \text{I}_2(\text{g})$ given that the concentrations of each species at equilibrium are as follows: $[\text{HI}] = 0.85 \text{ mol/L}$, $[\text{I}_2] = 0.60 \text{ mol/L}$, $[\text{H}_2] = 0.27 \text{ mol/L}$.

- A) 5.25 B) 0.22 C) 4.5 D) 0.19 E) 1.6×10^2

Chemical Equilibrium Quiz

Margaret G. Kimble



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