

Chemistry Benchmark II Study Guide

1. Write the formula for the following:
  - iron (II) oxide
  - calcium phosphate
  - ammonium phosphate
  - copper (III) bromide
  - nitrogen dioxide
  - titantium (IV) oxide
  - potassium bromide
  - magnesium hydroxide
  - xenon hexafluoride
2. What is a binary compound? What is a ternary compound?
3. Write formulas for:
  - hydroiodic acid
  - carbonic acid
  - hydrosulfuric acid
  - hydrochloric acid
  - hypochlorous acid
  - perchloric acid
  - phosphorous acid
  - chloric acid
  - sulfuric acid
  - chlorous acid
4. Name these:
  - NaCl
  - KI
  - CaBr
  - Rb<sub>2</sub>O
  - Sr<sub>2</sub>
  - CuCl
  - HgO
  - Fe<sub>2</sub>O<sub>3</sub>
  - MnO<sub>2</sub>
  - PbCl<sub>4</sub>
  - SnBr<sub>4</sub>
  - BF<sub>3</sub>
  - NO
  - N<sub>2</sub>O<sub>5</sub>
  - CCl<sub>4</sub>
  - P<sub>2</sub>O<sub>5</sub>
  - N<sub>2</sub>O<sub>3</sub>
5. Find the oxidation number assigned to:
  - S in H<sub>2</sub>SO<sub>4</sub>
  - P in PCl<sub>5</sub>
  - S in SO<sub>3</sub>
6. Name these:
  - NO<sub>2</sub>
  - N<sub>2</sub>O
  - SO<sub>3</sub>
7. How many valence electrons do elements have in the following groups?
  - alkali metals
  - alkaline earth metals
  - halogens
  - noble gases
8. Write the complete electron configuration, noble gas configuration, Lewis dot structure, and orbital notation for:
  - Sr
  - Br
  - Na
  - K
  - N
  - Cl
9. In each group, identify the atom with the largest atomic radius.
  - Sn, Xe, Rb, Sr
  - IRn, He, Xe, Kr
  - Pb, Ba, Cs, At
10. In each group, identify the atom with the highest ionization energy.
  - Cs, K, Li
  - Ba, Sr, Ca
  - I, Br, Cl
  - Mg, Si, S
11. Define the following: ionic bond, covalent bond, polar covalent bond, nonpolar covalent bond.
12. Define electronegativity and describe the trend across a period and down a group. What type of element has the highest electronegativity? What type has the lowest electronegativity?
13. Identify the type of bond in each:
  - K-N
  - Cs-O
  - C-N
  - O-F
  - Na-Cl
  - Cl-F
  - Cl-O
  - Ca-Cl
14. Who first arranged the periodic table in order of increasing atomic mass?
15. Draw Lewis structures for the following and predict the molecular shape.
  - H<sub>2</sub>S
  - PH<sub>3</sub>
  - O<sub>2</sub>
  - sulfate ion
  - SO<sub>3</sub>
  - BF<sub>3</sub>
16. Calculate the wavelength and energy of 96.6MHz radio waves.
17. What is the frequency and energy of electromagnetic radiation having a wavelength of  $3.33 \times 10^{-9}$ m?
18. Predict the ions that the following elements might form:
  - F
  - K
  - I
  - Ge
  - Fe (trick question)
  - P
  - Ag

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**Rose Arny**

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