

ANTACIDS

Introduction:

Chemists are often rather puzzled about how to check the claims that manufacturers make about their products. For example, there are no reproducible experimental techniques available for determining how fast a particular product cures a headache, relieves an upset stomach, or revives tired blood. The scientific approach is to try to duplicate as many conditions as possible and control the remaining variables one at a time.

Next to treating pain, one of the most frequent uses of over-the-counter medicines is the treatment of indigestion caused by an overproduction of stomach acid (HCl). The number of over-the-counter antacids is almost overwhelming, although the variation of ingredients is small. The most common antacids contain one or more of the following: sodium bicarbonate, calcium carbonate, magnesium carbonate, aluminum hydroxide, magnesium hydroxide, and magnesium trisilicate. Each of these compounds can neutralize stomach acid.

Table I – Antacid Reactions

Basic Ion	Compound	Reaction
bicarbonate	NaHCO_3 or KHCO_3	$\text{HCO}_3^- + \text{HCl} \leftrightarrow \text{Cl}^- + \text{H}_2\text{CO}_3^*$
hydroxide	$\text{Mg}(\text{OH})_2$, $\text{Al}(\text{OH})_3$, or $\text{NaAl}(\text{OH})_2\text{CO}_3$	$\text{OH}^- + \text{HCl} \leftrightarrow \text{Cl}^- + \text{H}_2\text{O}$
carbonate	CaCO_3 , MgCO_3 , or $\text{NaAl}(\text{OH})_2\text{CO}_3$	$\text{CO}_3^{2-} + 2 \text{HCl} \leftrightarrow 2 \text{Cl}^- + \text{H}_2\text{CO}_3^*$
trisilicate	$\text{Mg}_2\text{Si}_3\text{O}_8$	$\text{Si}_3\text{O}_8^{4-} + 4 \text{HCl} \leftrightarrow 4 \text{Cl}^- + 3 \text{SiO}_2 + 2 \text{H}_2\text{O}$
citrate	$\text{Na}_3\text{C}_6\text{H}_5\text{O}_7$	$\text{C}_6\text{H}_5\text{O}_7^{3-} + 3 \text{HCl} \leftrightarrow 3\text{Cl}^- + \text{H}_3\text{C}_6\text{H}_5\text{O}_7$
tartrate	$\text{Na}_2\text{C}_4\text{H}_4\text{O}_6$	$\text{C}_4\text{H}_4\text{O}_6^{2-} + 2 \text{HCl} \leftrightarrow 2 \text{Cl}^- + \text{H}_2\text{C}_4\text{H}_4\text{O}_6$

* carbonic acid is unstable and decomposes into water and carbon dioxide in solution

The various antacids have advantages and disadvantages. Sodium salts should be avoided by individuals on low-sodium diets. Magnesium salts may have constipating effects. Combinations of aluminum and magnesium salts are sometimes formulated in an attempt to prevent side effects.

Effervescent antacids utilize citric acid and sodium bicarbonate in a reaction to produce carbon dioxide bubbles. More than 85% of the bicarbonate of Alka-Seltzer® is consumed by the effervescent action. Do you think that it is possible to determine the acid neutralizing capability of an antacid by simply measuring the amount of a single chemical contained in the antacid?

* carbonic acid (H_2CO_3) is unstable and decomposes in solution, forming water and carbon dioxide

Chemistry Lab Antacids

Nicholas D. Priest, Frank Van de Vyver

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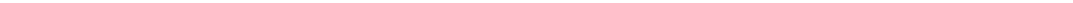
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