

Study Guide: Chapter 2 – Chemistry of Life

What are the components of an atom? What is the property of each component (charge and relative size)?

Nucleus, that contains protons (+, it is 1 Da = 1.7×10^{-26} grams) and neutrons (no electric charge, 1 Da). Around the nucleus there are also electrons (-, 0.0005 Da = 9×10^{-31} gram).

What is the atomic number, atomic mass and atomic weight?

- **Atomic number:** the number of protons in the nucleus, it is unique to each element
- **Atomic mass:** the total number of protons and neutrons in an atom (except hydrogen, because it doesn't have neutrons)
- **Atomic weight:** the ratio of the average mass per atom of the element to 1/12 of the mass of an atom of ^{12}C .

What are valent electrons? What is the octet rule? When is it complete? Is it the same for Hydrogen and Helium?

A valent electron is an electron in the outer shell of an atom.

An octet rule is the tendency of atoms to form stable molecules so that they have eight electrons in their outermost shell. (It is complete if they have eight electrons on their outermost shell)

What is an isotope? Does it have the same atomic mass? The same atomic number?

Different isotopes of the same element have the same number of protons but different numbers of neutrons. Isotopes are formed when atoms combine and/or release subatomic particles (decay).

They are atoms that have different atomic masses, which have the same atomic number.

Hydrogen and helium is an exception from the octet rule, because they attain stability, when two electrons occupy their single shell.

What is a half-life ($t_{1/2}$) with regards to radioactive decay?

The half-life of a radioactive isotope is the amount of time it takes for one-half of the radioactive isotope to decay.

When will an atom be positively charged? Negatively charged?

An atom is positively charged when it has more protons than electrons, in this case they are called cations. If there are more electrons than protons in an atom it's charged negatively and they are called anions.

What is the difference between polar and non-polar? Hydrophobic and hydrophilic?

Polar molecules have a uniform distribution of electron density. A nonpolar molecule results from an unequal distribution of electron density.

What are the forces that hold atoms together in a covalent bond, ionic bond, hydrogen bond, hydrophobic forces, van der Waals? Which bonds are the hardest to break?

- A covalent bond forms when two atoms attain stable electron numbers in their outermost shells by sharing one or more pairs of electrons.
- Ionic bonds form when an atom gains or loses one or more electrons to achieve stability.

Chemistry Of Life Study Guide

Marvin L. Hackert

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