



#### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
potassium-40 undergoes beta decay (electron)  
uranium-238 undergoes alpha decay

#### electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



#### NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

a. Sulfur hexafluoride  $\text{SF}_6$

b. Lithium nitride  $\text{Li}_3\text{N}$

c. Chromium (III) carbonate  $\text{Cr}_2(\text{CO}_3)_3$

d. Tin (II) chloride  $\text{SnCl}_2$

e. Ammonium acetate  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  (or  $\text{CH}_3\text{COONH}_4$ )

f. Mercury (I) chloride  $\text{Hg}_2\text{Cl}_2$

g. Potassium bromate  $\text{KBrO}_3$

h. Hydrosulfuric acid  $\text{H}_2\text{S(OH)}_2$

i. Chloric acid  $\text{HClO}_3$

j. Sulfurous acid  $\text{H}_2\text{SO}_3$

#### Name each of the following compounds.

a.  $\text{CuSO}_4$  Copper (II) sulfate

b.  $\text{AlF}_3$  aluminum fluoride

c.  $\text{HI}$  hydrogen iodide

d.  $\text{NO}$  nitrogen monoxide

e.  $\text{H}_2\text{Se}$  hydrogen selenide

f.  $\text{HNO}_3$  nitric acid

g.  $\text{HNO}_2$  nitrous acid

h.  $\text{NaHSO}_4$  sodium hydrogen sulfate

#### Half-life

Tritium ( $\text{H-3}$ ) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it takes for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25 \quad \boxed{61.5 \text{ yr}}$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} \boxed{12.5 \text{ mg}}$$

#### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{r} 63 \times .692 = 43.596 \rightarrow 43 \\ 65 \times .308 = 20.02 \rightarrow 20 \\ \hline 63.62 \rightarrow 64 \end{array}$$

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