

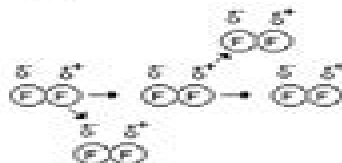
CHAPTER 11 PRACTICE QUESTIONS KEY

INSTRUCTORS: Do not release this key until noon 4/22/18.

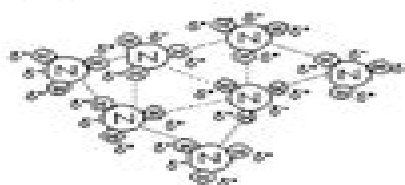
Use the following to answer questions 1-3:

Consider the representations below to answer the next three questions.

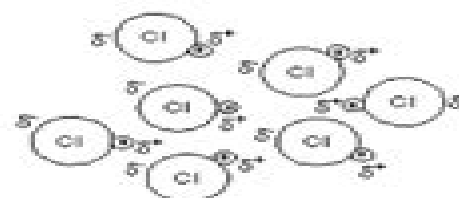
(I)



(II)



(III)



(III)

- How many of the following statements are correct concerning drawing I?
 - Each molecule induces a dipole onto the next molecule in close proximity.
 - The phenomenon shown is relatively weak and short-lived.
 - C_6H_{18} contains this type of interaction.
 - The forces that exist in this example are London dispersion forces.

a) 0 b) 1 c) 2 d) 3 e) 4
- Which drawing *best* represents the interactions in a sample of HF?

a) I
b) **II**
c) III
d) I, II
e) all of the above
- Which of the following statements are incorrect concerning drawing III?

a) Electrostatic interactions exist between the molecules.
b) The molecules find the best compromise between attraction and repulsion.
c) **These molecules exhibit ionic bonding.**
d) OCS exhibits this type of interaction.
e) Two of the above statements are incorrect.
- Which of the following is the correct order of boiling points for KNO_3 , CH_3OH , C_2H_6 , Ne?

a) $Ne < CH_3OH < C_2H_6 < KNO_3$
b) $KNO_3 < CH_3OH < C_2H_6 < Ne$
c) $Ne < C_2H_6 < KNO_3 < CH_3OH$
d) **$Ne < C_2H_6 < CH_3OH < KNO_3$**
e) $C_2H_6 < Ne < CH_3OH < KNO_3$
- What is the strongest intermolecular force (IMF) found in a sample of thousands of HBr molecules?

a. London b. **dipole dipole** c. nonpolar d. H bond force e. polar

Chemistry 11 Practice Exam With Answers

Darrell D. Ebbing, David Bookin



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Table of Contents Chemistry 11 Practice Exam With Answers

1. Understanding the eBook Chemistry 11 Practice Exam With Answers
 - The Rise of Digital Reading Chemistry 11 Practice Exam With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 11 Practice Exam With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 11 Practice Exam With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 11 Practice Exam With Answers
 - Personalized Recommendations

- Chemistry 11 Practice Exam With Answers User Reviews and Ratings
- Chemistry 11 Practice Exam With Answers and Bestseller Lists
- 5. Accessing Chemistry 11 Practice Exam With Answers Free and Paid eBooks
 - Chemistry 11 Practice Exam With Answers Public Domain eBooks
 - Chemistry 11 Practice Exam With Answers eBook Subscription Services
 - Chemistry 11 Practice Exam With Answers Budget-Friendly Options
- 6. Navigating Chemistry 11 Practice Exam With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry 11 Practice Exam With Answers Compatibility with Devices
 - Chemistry 11 Practice Exam With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry 11 Practice Exam With Answers
 - Highlighting and Note-Taking Chemistry 11 Practice Exam With Answers
 - Interactive Elements Chemistry 11 Practice Exam With Answers
- 8. Staying Engaged with Chemistry 11 Practice Exam With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry 11 Practice Exam With Answers
- 9. Balancing eBooks and Physical Books Chemistry 11 Practice Exam With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 11 Practice Exam With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry 11 Practice Exam With Answers
 - Setting Reading Goals Chemistry 11 Practice Exam With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry 11 Practice Exam With Answers
 - Fact-Checking eBook Content of Chemistry 11 Practice Exam With Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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