

Chem 12 Practice Solubility Test

1. Which combination of factors will affect the rate of the following reaction?



- A. Temperature and surface area only
- B. Temperature and concentration only
- C. Concentration and surface area only
- D. Temperature, concentration, and surface area

2. Consider the following reaction:



The rate of decomposition of the oxalate ion is increased by

- A. Adding NaOH
- B. Removing CO_2
- C. Adding a catalyst
- D. Decreasing the pressure

3. An equilibrium system shifts left when the temperature is increased. The forward rate is

- A. Exothermic and ΔH is positive.
- B. Exothermic and ΔH is negative.
- C. Endothermic and ΔH is positive.
- D. Endothermic and ΔH is negative.

4. The value of the K_{eq} can be changed by

- A. adding a catalyst.
- B. changing the temperature.
- C. changing the reactant concentration.
- D. changing the volume of the container

5. Consider the following equilibrium: $2\text{NOCl}_{(g)} \rightleftharpoons 2\text{NO}_{(g)} + \text{Cl}_{2(g)}$

In a 1.0 L container at equilibrium there are 1.0 mol NOCl, 0.70 mol NO and 0.40 mol Cl_2 . At constant temperature and volume, 0.10 mol NOCl is added. The concentrations in the "new" equilibrium in comparison to the concentrations in the "old" equilibrium are

	[NOCl]	[NO]	[Cl ₂]
A	new = old	new = old	new = old
B	new > old	new > old	new > old
C	new < old	new < old	new > old
D	new < old	new > old	new > old

Chemistry 12 Solubility Questions

Annelies Wilder-Smith



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