



Introduction to working as a Chemist.

Use a highlighter as you work through this booklet. ESL students, make vocab lists of the terminology that is defined or described here.

Assessment Statements

- Outline the Scientific Method**
- Identify independent, dependent & controlled variables.**
- Demonstrate correct and safe use of laboratory equipment.**
- Distinguish between quantitative and qualitative data**
 - > Qualitative is descriptive, not using numbers, e.g. properties, colour, gas production, sound, odor.
 - > Quantitative is mathematical, including
 - Raw data (as recorded)
 - Processed (after calculations)
- Distinguish between accuracy and precision**
 - > Accuracy describes the correctness of a set of measurements
 - > Precision describes the repeatability of a set of measurements
- State SI units of measurement for**
 - > Mass: grams (g)
 - > Length: Metres (m), millimeters (mm)
 - > Concentration: molarity (M)
 - > Temperature: Degrees Celsius (°C)
 - > Volume: Litres (L) or millilitres (mL)
 - > Density: grams per cubic centimeter (g/cm³)
- Conductivity: Siemens (S)**
- Acidity or alkalinity (pH)**
- Energy: Joules (J or kJ)**
 - Determine the uncertainty of digital and analog measuring tools
 - > Usually ± 1 of the smallest division
- Calculate the % error of a given or recorded set of values**
 - > $\frac{(\text{Measured} - \text{True})}{\text{True}} \times 100$
- State values to appropriate numbers of significant figures**
- Calculate values to appropriate numbers of significant figures**

Related Concepts:

- 01.1 Element symbols (Quia quiz: <http://www.quia.com/quiz/3246640.html>)
- 01.2 Accuracy, Precision, Error, Uncertainty (Quia quiz: <http://www.quia.com/quiz/3507626.html>)
- 01.3 SigDigs and Notation (Quia quiz: <http://www.quia.com/quiz/3504378.html>)

Command Terms: Be sure you understand what is being asked of you

Analyse Break down to identify key parts. Interpret information to reach conclusions.

Apply	Use knowledge & understanding in response to a real situation or given circumstances.
Calculate	Obtain a numerical answer, showing relevant stages in the working.
Define	Give the precise meaning of a word, phrase, concept or physical quantity.
Describe	Give a detailed account or picture of an event, pattern, process or pattern.
Design	Produce a plan, simulation or model.
Determine	Obtain the only possible answer.
Discuss	Offer a considered and balanced review that includes a range of arguments, factors or hypotheses. Opinions or conclusions should be presented clearly and supported by appropriate evidence.
Distinguish	Make clear the differences between two or more concepts or items.
Evaluate	Assess the implications and limitations; make judgments about the ideas, works, solutions or ideas in relation to selected criteria.
Explain	Give a detailed account including reasons or causes.
Identify	Provide an answer from a number of possibilities. Recognise and state briefly a distinguishing fact or feature.
Label	Add title, labels or brief explanation(s) to a diagram or graph.
List	Give a sequence of answers with no explanation.
Outline	Give a brief account.
Recall	Remember or recognize from prior learning experiences.
State	Give a specific name, value or other brief answer without explanation or calculation.
Suggest	Propose a solution, hypothesis or other possible answer.

Chemistry 153 Lab Manual Answers

Jerome Alexander

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A Guide to the Literature of Chemistry Evan Jay Crane, Austin McDowell Patterson, 1927 A Laboratory Manual of Analytical Methods of Protein Chemistry, Including Polypeptides Peter Alexander, Richard Joseph Block, Harold Palmer Lundgren, 1960 In the last fifteen years there has been a revolution in the techniques available for the analysis and isolation of proteins Every time a new technique has been introduced numerous papers have appeared describing modifications to it and the research worker who wishes to employ these methods is faced with a very serious problem in deciding which particular variant to use These volumes are intended to provide the fullest practical detail so that any scientist can follow the procedure by using this book alone and without having recourse to the original literature The techniques which are described in full are ones in which all the authors have had first hand experience and the descriptions contain those small but important points which save so much time In the first volume separation and isolation procedures are discussed the second concerns its analysis and reactivity and the third volume with the measurement of the macromolecular properties of proteins

Practical Physiological Chemistry Philip Bovier Hawk, Olaf Bergeim, 1926 **A Laboratory Manual of Biological Chemistry** Adam Arthur Christman, 1954 The Chemical News, 1877 Clinical Chemistry Laboratory Manual Ruthanne Hyduke, Lucy J. Wall, 1989 Photometric Chemical Analysis John Howe Yoe, 1928 The Elements of Qualitative Chemical Analysis Julius Stieglitz, 1911 The Elements of qualitative chemical analysis v. 1 Julius Stieglitz, 1911 Chemical News and Journal of Physical Science, 1895 **Colloid Chemistry, Theoretical and Applied: Theory and methods** Jerome Alexander, 1926 The Chemical News and Journal of Industrial Science, 1921 The Chemical News and Journal of Physical Science, 1897 Chemical News and Journal of Industrial Science, 1921 Chemical News and Journal of Physical Science William Crookes, 1921 Chemical Engineering Catalog, 1922

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