

Chemistry

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(Chapter 11) (Alcohols Phenols and Ethers) (Exercises)

(Class 12)



Question 11.15:

Explain why is *ortho* nitrophenol more acidic than *ortho* methoxyphenol?

Answer 11.15:



o – Nitrophenol



o – Methoxyphenol

The nitro-group is an electron-withdrawing group. The presence of this group in the *ortho* position decreases the electron density in the O-H bond. As a result, it is easier to lose a proton. Also, the *o*-nitrophenoxide ion formed after the loss of protons is stabilized by resonance. Hence, *ortho* nitrophenol is a stronger acid.

On the other hand, methoxy group is an electron-releasing group. Thus, it increases the electron density in the O-H bond and hence, the proton cannot be given out easily.

For this reason, *ortho*-nitrophenol is more acidic than *ortho*-methoxyphenol.

Question 11.16:

Explain how does the –OH group attached to a carbon of benzene ring activate it towards electrophilic substitution?

Answer 11.16:

The -OH group is an electron-donating group. Thus, it increases the electron density in the benzene ring as shown in the given resonance structure of phenol.



As a result, the benzene ring is activated towards electrophilic substitution.

Question 11.17:

Give equations of the following reactions:

- Oxidation of propan -1-ol with alkaline KMnO₄ solution.
- Bromine in C₂S with phenol.
- Dilute HNO₃ with phenol.
- Treating phenol with chloroform in presence of aqueous NaOH.

Chemistry Chapter 11 Solutions Manual

John McMurry, Susan McMurry

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