

Assessment**Chapter Test B****Chapter: Chemical Equations and Reactions**

Directions: In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

1. The generation of a relatively unstable, short-lived compound by an alpha-disintegration reaction is called the **transmutation** of an atom.

- radioactive
- radioisotope
- radioactive reaction
- radioactive

2. The products of a chemical equation, in the presence of the catalyst, are:

- new elements
- new isotopes
- new atoms
- new molecules

3. In a chemical equation, the greatest (biggest) molecule that can be split is the **reactant**.

- product
- reactant
- new solid
- new solution

4. The tendency that a radioactive isotope has to decay back into a more stable element is called **radioactivity**.

- radioactive elements
- radioactive substances
- radioactive substances, like helium, gas
- radioactive substances that decay to other radioactive heavier elements

5. The characteristic for a radioactive isotope:

- isotopes that contain odd numbers of protons and neutrons
- shorter the半-life of an isotope and more radioactive it is
- more the number isotope of other substances
- shorter the half-life radioactive that reacts faster

6. If one of ^{238}U isotope loses an alpha, it becomes radioactive isotope, ^{234}Th .

- ^{238}U will radioactive isotope of ^{234}Th is radioactive
- ^{238}U will radioactive isotope of ^{234}Th is radioactive
- ^{238}U will become radioactive, really radioactive
- ^{238}U will react with a cold neutron

Chemistry Chapter 8 Study Guide

John A. Olmsted, Gregory M. Williams

Chemistry Chapter 8 Study Guide:

A Self-study Guide to the Principles of Organic Chemistry Jiben Roy, 2013 A Self Study Guide to the Principles of Organic Chemistry Key Concepts Reaction Mechanisms and Practice Questions for the Beginner will help students new to organic chemistry grasp the key concepts of the subject quickly and easily as well as build a strong foundation for future study Starting with the definition of atom the author explains molecules electronic configuration bonding hydrocarbons polar reaction mechanisms stereochemistry reaction varieties organic spectroscopy aromaticity and aromatic reactions biomolecules organic polymers and a synthetic approach to organic compounds The over one hundred diagrams and charts contained in this volume will help students visualize the structures and bonds as they read the text and make the logic of organic chemistry clear and easily understood Each chapter ends with a list of frequently asked questions and answers followed by additional practice problems Answers are included in the Appendix

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