

CHAPTER

02

SOLUTION

- Solution :** A homogeneous mixture of two or more substances is known as solution.
- Solute :** The substance present in smaller amount in a solution is called solute.
- Solvent :** The substance present in larger amount in a solution is called solvent.
- Binary solution :** Those solutions which contain two components are called binary solutions, e.g., salt solution, benzene and toluene.
- Ternary solutions :** Those solutions which contain three components are called ternary solutions, e.g., ethanol + water + acetic acid.
- Aqueous solution :** When solute is dissolved in water, it is called aqueous solution, e.g., sugar solution, ethanol in water.
- Non-aqueous solution :** When solute is dissolved in solvent other than water, it is called non-aqueous solution, e.g., iodine dissolved in alcohol (Tincture of iodine).
- Weight Fraction** =
$$\frac{\text{weight of solute}}{\text{weight of solution}}$$
- Weight percentage**
=
$$\frac{\text{weight of solute}}{\text{weight of solution}} \times 100$$

Weight of solute present in 100 grams of solution is called weight percentage
- Volume of solute present in 100 ml of the solution is called volume percentage

- Volume percentage**
=
$$\frac{\text{Volume of solute}}{\text{Volume of solution}} \times 100$$
- Volume Fraction** =
$$\frac{\text{Volume of solute}}{\text{Volume of solution}}$$
- Molarity (M)**
=
$$\frac{\text{No of moles of solute}}{\text{Volume of solution in liters}}$$

(or)
$$M = \frac{W}{GMW} \times \frac{1}{V(lit)}$$
- Molarity** =
$$\frac{\text{No of moles of solute}}{1000} \times \frac{1000}{\text{Volume of solution in ml}}$$
- Molarity** =
$$\frac{\text{Weight of solute}}{\text{GMWt of solute}} \times \frac{1000}{V \text{ in ml}}$$

(or)
$$M = \frac{W}{GMW} \times \frac{1000}{V \text{ (ml)}}$$
- Weight of solute** =
$$\frac{M \times GMW \times V \text{ in ml}}{1000}$$
- Molarity for dilute solutions** :
$$M_1 V_1 = M_2 V_2$$
- For neutralisation reactions** :
$$\frac{M_1 V_1}{n_1} = \frac{M_2 V_2}{n_2}$$
- For mixed solution molarity**
=
$$\frac{M_1 V_1 + M_2 V_2 + \dots}{V_1 + V_2 + \dots}$$

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Petr Beneš, Vladimír Majer

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