

**Assessment #3 Review**

You must show all work for credit.

Due Date: \_\_\_\_\_ (No late packets will be accepted)

Name: \_\_\_\_\_

**I. Balance Equations**

Balance the equations

1.  $Mg + 2H_2O \rightarrow Mg(OH)_2 + H_2$
2.  $Mg + 2NaBr \rightarrow MgBr_2 + 2Na$
3.  $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$
4.  $4P + 5O_2 \rightarrow 2P_2O_5$
5.  $Fe_2(SO_4)_3 + 6KOH \rightarrow 3K_2SO_4 + 2Fe(OH)_3$
6.  $Fe_2O_3 + 3H_2 \rightarrow Fe + 3H_2O$

7. How many moles of hydrogen gas react with 1 mole of  $Fe_2O_3$  in reaction #6? 3

Write the balanced equation for problems # 8 and 9:

8. Hydrogen gas and nitrogen monoxide combine to form water and nitrogen gas



9. When sodium phosphate and calcium chloride are mixed they form calcium phosphate and sodium chloride



10. Using the following balanced equation, prove (or explain) the Law of Conservation of Mass.



$\frac{1}{2}N$   $\frac{3}{2}H_2$   $\frac{2}{2}NH_3$  left = right  
 mass = mass gains weight loses weight

11. Using the balanced equation shown in problem #10, explain what the coefficients represent.

1 3 2

12. Using the balanced equation shown in problem #10, list the products of the reaction.



13. Using the balanced equation shown in problem #10, list the reactants of the reaction.

**II. Types of Reactions**

Directions: Identify the type of reaction.

S = synthesis

SD = single displacement

C = combustion

D = decomposition

DD = double displacement



# Chemistry Final Exam Just Guide Answers

**Don E. Webster**

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