

Match the notation below with the atom described in questions 1-4:

- (A) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}} \uparrow$
- (B) $1s \underline{\hspace{1cm}} \downarrow \uparrow 2s \underline{\hspace{1cm}} \downarrow \uparrow$
- (C) $1s \underline{\hspace{1cm}} \downarrow \uparrow 2s \underline{\hspace{1cm}} \downarrow \uparrow 2p \underline{\hspace{1cm}} \uparrow \underline{\hspace{1cm}} \uparrow \underline{\hspace{1cm}}$
- (D) $1s \underline{\hspace{1cm}} \downarrow \uparrow 2s \underline{\hspace{1cm}} \downarrow \uparrow 2p \underline{\hspace{1cm}} \downarrow \uparrow \underline{\hspace{1cm}} \downarrow \uparrow \underline{\hspace{1cm}} \downarrow \uparrow$
- (E) $[Ar] 4s \underline{\hspace{1cm}} \downarrow \uparrow 3d \underline{\hspace{1cm}} \downarrow \uparrow \underline{\hspace{1cm}} \uparrow \underline{\hspace{1cm}} \uparrow \underline{\hspace{1cm}} \uparrow$

- 1) Represents an atom that is in an excited state
 2) Represents an atom that is a noble gas
 3) Represents an atom that is a transition metal
 4) Represents an atom of an alkali earth metal

Questions 5-8 refer to the following descriptions of bonding in different types of solids.

- a) Lattice of positive and negative ions held together by electrostatic forces
 b) Closely packed lattice with delocalized electrons throughout giving ability to conduct electricity and permitting ductility
 c) Strong single covalent bonds making up a molecule with a low molecular mass
 d) Double and/or triple bonds are used to hold atoms together
 e) Expanded octet of electrons on the central atom
- 5) Potassium Chloride, $KCl(s)$
 6) Gold, $Au(s)$
 7) Methane, $CH_4(g)$
 8) Sulfur Hexafluoride, $SF_6(g)$

Questions 9-12 refer to the reactions represented below. Answers may be used more than once.

- a) $H_2SeO_4(aq) + 2Cl^-(aq) + 2H^+(aq) \longrightarrow H_2SeO_3(aq) + Cl^-_2(aq) + H_2O(l)$
 b) $C_2H_6(g) + O_2(g) \longrightarrow CO_2(g) + H_2O(l)$
 c) $3I^-(aq) + 6OH^-(aq) \longrightarrow 5I^-(aq) + IO_3^-(aq) + 3H_2O(l)$
 d) $Ba^{2+}(aq) + SO_4^{2-}(aq) \longrightarrow BaSO_4(s)$
 e) $Ag(s) + 2Cl^-(aq) \longrightarrow AgCl(s)$

- 9) A reaction with both a weak and a strong acid as reactants
 10) A precipitation reaction in which Ca^{2+} could be a spectator ion
 11) A reaction in which the same reactant undergoes both oxidation and reduction
 12) A combustion reaction

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