

Match the notation below with the atom described in questions 1-4:



- 1) Represents an atom that is in an excited state
- 2) Represents an atom that is a noble gas
- 3) Represents an atom that is a transition metal
- 4) Represents an atom of an alkali earth metal

Questions 5-8 refer to the following descriptions of bonding in different types of solids.

- a) Lattice of positive and negative ions held together by electrostatic forces
 - b) Closely packed lattice with delocalized electrons throughout giving ability to conduct electricity and permitting ductility
 - c) Strong single covalent bonds making up a molecule with a low molecular mass
 - d) Double and/or triple bonds are used to hold atoms together
 - e) Expanded octet of electrons on the central atom
- 5) Potassium Chloride, $\text{KCl}_{(s)}$
 - 6) Gold, $\text{Au}_{(s)}$
 - 7) Methane, $\text{CH}_{4(s)}$
 - 8) Sulfur Hexafluoride, $\text{SF}_{6(g)}$

Questions 9-12 refer to the reactions represented below. Answers may be used more than once.



- 9) A reaction with both a weak and a strong acid as reactants
- 10) A precipitation reaction in which Ca^{2+} could be a spectator ion
- 11) A reaction in which the same reactant undergoes both oxidation and reduction
- 12) A combustion reaction

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