

Multiple Choice

Refer to the set of graphs below when answering the next two questions:

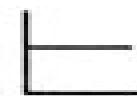
Graph #1



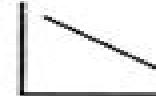
Graph #2



Graph #3



Graph #4



B 1. Which of the graphs above is the best qualitative graph of pressure versus volume of nitrogen at constant temperature and number of moles? → INVERSE

- A) Graph #1
- B) Graph #2
- C) Graph #3
- D) Graph #4

A 2. Which of the graphs above is the best qualitative graph of Kelvin temperature versus pressure of an ideal gas at constant volume and number of moles? → DIRECT

- A) Graph #1
- B) Graph #2
- C) Graph #3
- D) Graph #4

D 3. Use the kinetic molecular theory of gases to predict what would happen to a closed sample of a gas whose temperature increased while its volume decreased. → more ENERGY, more COLLISIONS

- A) The average kinetic energy of the molecules of the gas would decrease.
- B) Its pressure would decrease.
- C) The number of moles of gas would decrease.
- D) Its pressure would increase.

B 4. A hydrocarbon gas with the empirical formula CH_2 has a density of 1.3 g/L at 0°C and 1.00 atm. A possible formula for the hydrocarbon is:

- A) CH_2
- B) C_2H_4
- C) C_3H_6
- D) C_4H_8
- E) C_5H_{10}

~~_____~~

$$1.3 \text{ g/L} \times 22.4 \text{ L/mol} \approx 30 \text{ g/mol}$$

Chemistry Gases Study Guide Answers

Xiaolong Qi

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