

## Chemistry Exam 1 Study Guide (Chapters 1-4)

### Chapter 1: Matter and Energy

#### 1.1 Matter and its Classification:

- Matter- anything that occupies space and has mass.
- Mass- a measure of the quantity of matter.
- Pure Substance- matter that has the same chemical composition no matter what its origin.
  - Elements and Compounds
- Mixture- a combination of two or more substances that can vary in composition.
  - Separation of mixtures is **physical** using procedures such as **grinding, dissolving, or filtering**.
- Homogeneous Mixture- a combination of two or more substances that has uniform composition.
  - Solution- a homogeneous mixture of two or more substances uniformly dispersed at a molecular or ionic level.
- Heterogeneous Mixture- a combination of two or more substances that is not uniform throughout.
- Element- a substance that cannot be broken down into simpler substances even by a chemical reaction.
  - Metal- an element characterized by luster and the ability to conduct electricity. **Ex: Copper(Cu), Aluminum(Al), Iron(Fe)**
  - Nonmetal- an element that usually has a dull appearance and is a poor conductor of electricity. **Ex: Carbon(C), Chlorine(Cl), Sulfur(S)**
- Compound- a substance composed of two or more elements combined in definite proportions.
- Chemical Formula- describes the composition of a compound, using the symbols for the elements that make up the compound.
- Atom- the smallest unit of an element that has the chemical properties of that element.
- Molecule- two or more atoms bound together in a discrete arrangement.
- Aqueous Solution- a homogeneous mixture of two or more substances in which the solvent is water.

#### Physical States of Matter

Solid	Liquid	Gas
fixed shape	shape of container (may or may not fill it)	shape of container (fills it)
its own volume	its own volume	volume of container
no volume change under pressure	slight volume change under pressure	large volume change under pressure
particles are fixed in place and tend to be in a regular (crystalline) array	particles are randomly arranged and free to move about until they bump into one another	particles are widely separated and move independently of one another

# Chemistry Matter And Change Course Planning Guide

**R Sandford**

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