

Chemistry

BASIC CONCEPTS

- The percentage of oxygen in NaOH is
A. 40 B. 60
C. 8 D. 10
- What quantity of limestone (CaCO_3) on heating will give 56 kg of CaO?
A. 1000 kg B. 56 kg
C. 44 kg D. 100 kg
- Which of the following has least mass?
A. 2 gram atom of nitrogen
B. 3×10^{23} atoms of C
C. 1 mole of S
D. 7.0 g of Ag.
- 1 mole of CH_4 contains
A. 6.02×10^{23} atoms of H
B. 4 g-atom of hydrogen
C. 1.81×10^{23} molecules of CH_4
D. 3.0 g of carbon
- How many moles of Helium gas occupy 22.4 L at 0°C at 1 atm. Pressure?
A. 0.11 B. 0.90
C. 1.0 D. 1.11
- The number of atoms contained in 11.2 L of SO_2 at S.T.P. are
A. $3/2 \times 6.02 \times 10^{23}$
B. $2 \times 6.02 \times 10^{23}$
C. 6.02×10^{23}
D. $4 \times 6.02 \times 10^{23}$
- The Number of oxygen atoms in 4.4 g of CO_2 is approximately

- A. 1.2×10^{23} B. 6×10^{22}
C. 6×10^{23} D. 12×10^{23}

- The total number of protons in 10 g of calcium carbonate is ($N_0 = 6.023 \times 10^{23}$)
A. 1.5057×10^{24}
B. 2.0478×10^{24}
C. 3.0115×10^{24}
D. 4.0956×10^{24}
- 2 g of oxygen contains number of atoms equal to that in
A. 0.5 g of hydrogen
B. 4 g of sulphur
C. 7 g of nitrogen
D. 2.3 g of sodium
- If N_A is Avogadro's number, then number of valence electrons in 4.2 g of nitride ions N^{3-} is
A. $2.4 N_A$ B. $4.2 N_A$
C. $1.6 N_A$ D. $3.2 N_A$

ANSWERS

1	A	6	A
2	D	7	A
3	B	8	C
4	B	9	B
5	C	10	A

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