

Study Guide: Unit #4 Mixtures and Solutions

1. Key Terms

Chemistry
Matter
Heterogeneous Mixture
Homogeneous Mixture
Pure Substance
Compounds
Solvent
Solubility
Insoluble
Qualitative
Dilute
Saturated
Supersaturated

Properties
Mixture
Mechanical Mixture
Solution
Elements
Solute
Dissolve
Soluble
Concentration
Concentrated
Quantitative
Unsaturated

1. Know the Particle Theory of matter
2. How can you tell the difference between a homogeneous and a heterogeneous mixture (by the unaided eye and in the lab)
3. Examples of mixtures that could be both homogeneous and heterogeneous
4. How is a pure substance different than a mixture
5. Use the particle theory of matter to explain the difference between a pure substance and a mixture
6. Two types of pure substances
7. Examples of elements and compounds
8. Identify the solute and solvent in a solution
9. Examples of different solutions with solutes and solvents of different states (see table 8.1 page 255)
10. Explain using the particle theory of matter why some substances dissolve and others do not
11. Explain what g/L means.
12. Determine the concentration in g/L of a solution when given the amount of solute in grams and the volume of solvent in either liters or milliliters
13. Determine when a situation is either qualitative or quantitative
14. Determine if a solution is dilute, concentrated, unsaturated, saturated or supersaturated when viewing a diagram or given a description.

Chemistry Mixtures And Solutions Study Guide

E Staff

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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