

Acid/Base indicator: \_\_\_\_\_ that changes \_\_\_\_\_ in the presence of an \_\_\_\_\_ or a \_\_\_\_\_.

Solution	Litmus paper (R→B, B→R, or NC)	Phenolphthalein (color or NC)	Bromothymol Blue (color or NC)	Cabbage Juice (color or NC)
HNO <sub>3</sub>				
NaOH				
KOH				
H <sub>2</sub> SO <sub>4</sub>				

Conclusion Questions:

1. Litmus turns \_\_\_\_\_ in an acid and \_\_\_\_\_ in a base.
2. Phenolphthalein (phth) turns \_\_\_\_\_ in an acid and \_\_\_\_\_ in a base.
3. Bromothymol blue turns \_\_\_\_\_ in an acid and \_\_\_\_\_ in a base.
4. Cabbage juice turns \_\_\_\_\_ in an acid and \_\_\_\_\_ in a base.

Strong Acids: dissociate \_\_\_\_\_ in \_\_\_\_\_ (ex: \_\_\_\_\_)  
 Weak Acids: dissociate \_\_\_\_\_ in \_\_\_\_\_  
 (ex: \_\_\_\_\_ or \_\_\_\_\_)

Strong Bases: dissociate \_\_\_\_\_ in \_\_\_\_\_ (ex: \_\_\_\_\_)  
 Weak Bases: dissociate \_\_\_\_\_ in \_\_\_\_\_  
 (ex: \_\_\_\_\_ or \_\_\_\_\_)

pH = \_\_\_\_\_

0 \_\_\_\_\_ 7 \_\_\_\_\_ 14

Determine the pH of a solution of HCl that has a molarity of  $1 \times 10^{-4} \text{ M}$ .

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**SA Adler**



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