



CHEMISTRY

PERIODIC TABLE OF THE ELEMENTS

PERIODIC TRENDS

- Atomic Size**
 - Increases from top-left to bottom-right.
 - Increases down a group.
 - Increases across a period from left to right.
- Ionization Energy**
 - Increases from bottom-right to top-left.
 - Increases across a period from left to right.
 - Increases up a group.
- Electronegativity**
 - Increases from bottom-right to top-left.
 - Increases across a period from left to right.
 - Increases up a group.

ATOMIC STRUCTURE AND THE PERIODIC TABLE

- Atomic Spectra**
 - Discrete lines of light.
 - Each element has a unique spectrum.
- Quantum Numbers**
 - Principal quantum number (n): size of orbital.
 - Angular momentum quantum number (l): shape of orbital.
 - Magnetic quantum number (m_l): orientation of orbital.
 - Spin quantum number (m_s): spin of electron.
- Wave Mechanical Model**
 - Electrons are described as waves.
 - Wave function (ψ) describes the probability of finding an electron.
 - Schrodinger equation: $\nabla^2 \psi + 2m(E - V)\psi = 0$

Quantitative Chemistry: Molar Mass and Avogadro's Law



Light elements up to 20 were considered up to 1913. After that, the periodic table was extended to include elements up to 118.

Electronic Structure of Atoms

- Electrons are arranged in shells.
- Each shell can hold a maximum of $2n^2$ electrons.
- Electrons fill shells from lowest energy to highest energy.
- Order of filling: 1s, 2s, 2p, 3s, 3p, 4s, 3d, 4p, 5s, 4d, 5p, 6s, 4f, 5d, 6p, 7s, 5f, 6d, 7p.



The Wave Mechanical Model

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Properties

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- Wave function (ψ) describes the probability of finding an electron.
- Schrodinger equation: $\nabla^2 \psi + 2m(E - V)\psi = 0$

Periodic Trends

- Atomic size increases from top-left to bottom-right.
- Ionization energy increases from bottom-right to top-left.
- Electronegativity increases from bottom-right to top-left.

Bohr's Model

- Electrons move in fixed orbits.
- Energy levels are quantized.
- Electrons can jump between energy levels by absorbing or emitting light.

Chemistry Quickstudy Reference Guides Academic

Mark Jackson



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Chemistry BarCharts Inc,2012-05-31 **Chemistry** Mark D. Jackson,BarCharts, Inc,2001-01-24 This 4 page laminated guide is essential for students or individual looking for Chemistry knowledge The QuickStudy Chemistry guide includes information on atomic structure nuclear chemistry atomic quantum numbers orbitals electronic atoms types of matter reactions physical processes nomenclature stoichiometry chemical interactions bonding models gasesthermodynamics theories and much more

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