

Chemistry Spring final exam review questions with answers below

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- _____ 1. A 1.0 M solution of hydrochloric acid conducts electricity better than a 1.0 M solution of acetic acid. Which of the following is the best explanation of this?
- HCl is a strong acid while $\text{HC}_2\text{H}_3\text{O}_2$ is weak.
 - The solutions are not the same pH.
 - Acetic acid has a higher molecular weight.
 - Acetic acid has more ions in one formula unit.
- _____ 2. The standard molar heat of vaporization for water is 40.79 kJ/mol. How much energy would be required to vaporize 94.0 g of water that started at 100°C and 1 atm pressure?
- 0.128 kJ
 - 7.81 kJ
 - 213 kJ
 - 3800 kJ
- _____ 3. What are the acids in the following equilibrium reaction?
- $$\text{CN}^-(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{HCN}(\text{aq}) + \text{OH}^-(\text{aq})$$
- CN^- , OH^-
 - H_2O , OH^-
 - CN^- , H_2O
 - H_2O , HCN
- _____ 4. If 4 moles of gas are added to a container that already holds 1 mole of gas, how will the pressure change inside the container?
- The pressure will be five times higher.
 - The pressure will be four times higher.
 - The pressure will decrease by one fifth.
 - The pressure will not change.
- _____ 5. A sample of gas occupies 17 mL at -112°C. What volume does the sample occupy at 70°C?
- 8.0mL
 - 27 mL
 - 36mL
 - 10.6 mL
- _____ 6. If the concentration for a solution is 4% by mass and the mass of the solution is 200 g, what is the mass of solute in solution?
- 8.0 g
 - 50 g
 - 800 g
 - 80 g
- _____ 7. The volume of 4.00M HCl that can be made from 3.19 mL of 6.80M HCl is_____.
- 281 mL
 - 8.52 mL
 - 5.42 mL
 - 1.87 mL
- _____ 8. What mass of sucrose (342 g/mol, $\text{C}_{12}\text{H}_{22}\text{O}_{11}$) is needed to make 400.0 mL of a 0.300M solution?
- 17.1 g
 - 34.2 g
 - 41.0 g
 - 28.5 g
- _____ 9. Sodium reacts with water according to the equation below. If 200. grams of sodium react, what volume of hydrogen gas is produced at 298K and 1.10 atm?
- $$2 \text{Na}(\text{s}) + 2 \text{H}_2\text{O}(\text{l}) \rightarrow 2 \text{NaOH}(\text{aq}) + \text{H}_2(\text{g})$$
- 96.8 L
 - 48.4 L
 - 193.6 L
 - 2046 L
- _____ 10. Which of the changes listed below would shift the following reaction to the left?
- $$4 \text{HCl}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2 \text{Cl}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{g})$$
- addition of Cl_2
 - removal of Cl_2

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JS Bruner



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