



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol \rightleftharpoons .

<p>What is Equilibrium?</p> <p>Equilibrium is a state of balance between two or more opposing processes.</p> <p>In a chemical reaction, it is the state where the concentrations of the reactants and products remain constant over time.</p> <p>Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the rates of the forward and reverse reactions are equal.</p>	<p>Equilibrium Constants</p> <p>K_c (Concentration): $K_c = \frac{[Products]}{[Reactants]}$</p> <p>$K_p$ (Pressure): $K_p = \frac{P_{Products}}{P_{Reactants}}$</p> <p>$K_a$ (Acid Dissociation): $K_a = \frac{[H^+][A^-]}{[HA]}$</p> <p>$K_b$ (Base Dissociation): $K_b = \frac{[OH^-][BH^+]}{[B]}$</p>													
<p>Equilibrium Constant, K_c</p> <p>For a general reaction: $aA + bB \rightleftharpoons cC + dD$</p> <p>The equilibrium constant K_c is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ <p>Where $[A]$, $[B]$, $[C]$, and $[D]$ are the concentrations of the reactants and products at equilibrium.</p>	<p>Equilibrium Constant, K_p</p> <p>For a general reaction: $aA + bB \rightleftharpoons cC + dD$</p> <p>The equilibrium constant K_p is given by:</p> $K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$ <p>Where P_A, P_B, P_C, and P_D are the partial pressures of the reactants and products at equilibrium.</p>													
<p>Equilibrium Constant, K_a</p> <p>For a weak acid HA dissociating in water:</p> $HA \rightleftharpoons H^+ + A^-$ <p>The equilibrium constant K_a is given by:</p> $K_a = \frac{[H^+][A^-]}{[HA]}$	<p>Equilibrium Constant, K_b</p> <p>For a weak base B reacting with water:</p> $B + H_2O \rightleftharpoons BH^+ + OH^-$ <p>The equilibrium constant K_b is given by:</p> $K_b = \frac{[BH^+][OH^-]}{[B]}$													
<p>Le Chatelier's Principle: The effect of changing conditions on equilibrium</p> <table><tr><th>Condition</th><th>Effect on Equilibrium</th><th>Example</th></tr><tr><td>Concentration</td><td><ul style="list-style-type: none">Increasing the concentration of a reactant shifts the equilibrium to the right (towards products).Increasing the concentration of a product shifts the equilibrium to the left (towards reactants).</td><td>$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$<p>Increasing $[N_2]$ or $[H_2]$ increases $[NH_3]$.</p></td></tr><tr><td>Pressure</td><td><ul style="list-style-type: none">Increasing the pressure (by decreasing volume) shifts the equilibrium towards the side with fewer moles of gas.</td><td>$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$<p>Increasing pressure shifts the equilibrium to the right (fewer moles of gas).</p></td></tr><tr><td>Temperature</td><td><ul style="list-style-type: none">For an exothermic reaction, increasing the temperature shifts the equilibrium to the left (towards reactants).For an endothermic reaction, increasing the temperature shifts the equilibrium to the right (towards products).</td><td>$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g) \quad \Delta H = -92 \text{ kJ/mol}$<p>Increasing temperature shifts the equilibrium to the left.</p></td></tr></table>			Condition	Effect on Equilibrium	Example	Concentration	<ul style="list-style-type: none">Increasing the concentration of a reactant shifts the equilibrium to the right (towards products).Increasing the concentration of a product shifts the equilibrium to the left (towards reactants).	$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ <p>Increasing $[N_2]$ or $[H_2]$ increases $[NH_3]$.</p>	Pressure	<ul style="list-style-type: none">Increasing the pressure (by decreasing volume) shifts the equilibrium towards the side with fewer moles of gas.	$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ <p>Increasing pressure shifts the equilibrium to the right (fewer moles of gas).</p>	Temperature	<ul style="list-style-type: none">For an exothermic reaction, increasing the temperature shifts the equilibrium to the left (towards reactants).For an endothermic reaction, increasing the temperature shifts the equilibrium to the right (towards products).	$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g) \quad \Delta H = -92 \text{ kJ/mol}$ <p>Increasing temperature shifts the equilibrium to the left.</p>
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Chemistry Study Guide Answers Chemical Equilibrium

Thomas Elliott Taylor



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