

Stoichiometry

Section 11.1 Defining Stoichiometry

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Practice Problems

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1. Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is observed.



1 molecule N_2 + 3 molecules

$\text{H}_2 \rightarrow$ 2 molecules NH_3

1 mole N_2 + 3 moles $\text{H}_2 \rightarrow$ 2 moles NH_3

Mass:

N_2 : (2 mol)(14.007 g/mol) = 28.014 g

3H_2 : (6 mol)(1.008 g/mol) = 6.048 g

2NH_3 : (2 mol)(14.007 g/mol) +
(6 mol)(1.008 g/mol) = 34.062 g

28.014 g N_2 + 6.048 g $\text{H}_2 \rightarrow$ 34.062 g NH_3

34.062 g reactants = 34.062 g products



1 molecule HCl + 1 formula unit $\text{KOH} \rightarrow$

1 formula unit KCl + 1 molecule H_2O

1 mole HCl + 1 mole $\text{KOH} \rightarrow$

1 mole KCl + 1 mole H_2O

Mass:

HCl : (1 mol)(1.008 g/mol) +
(1 mol)(35.453 g/mol) = 36.461 g

KOH : (1 mol)(39.098 g/mol) +
(1 mol)(15.999 g/mol) + (1 mol)(1.008 g/mol)
= 56.105 g

KCl : (1 mol)(39.098 g/mol) +
(1 mol)(35.453 g/mol) = 74.551 g

H_2O : (2 mol)(1.008 g/mol) +
(1 mol)(15.999 g/mol) = 18.015 g

36.461 g HCl + 56.105 g $\text{KOH} \rightarrow$
74.551 g KCl + 18.015 g H_2O

92.566 g reactants = 92.566 g products



2 atoms Mg + 1 molecule

$\text{O}_2 \rightarrow$ 2 formula units MgO

2 moles Mg + 1 mole $\text{O}_2 \rightarrow$ 2 moles MgO

Mass:

2Mg : (2 mol)(24.305 g/mol) = 48.610 g

O_2 : (2 mol)(15.999 g/mol) = 31.998 g

2MgO : (2 mol)(24.305 g/mol) +
(2 mol)(15.999 g/mol) = 80.608 g

48.610 g Mg + 31.998 g $\text{O}_2 \rightarrow$ 80.608 g MgO

80.608 g reactants = 80.608 g products

2. **Challenge** For each of the following, balance the chemical equation; interpret the equation in terms of particles, moles, and mass; and show that the law of conservation of mass is observed.



2 atoms Na + 2 molecules $\text{H}_2\text{O} \rightarrow$

2 formula units NaOH + 1 molecule H_2

2 mol Na + 2 mol $\text{H}_2\text{O} \rightarrow$

2 mol NaOH + 1 mol H_2

Mass:

2Na : (2 mol)(22.990 g/mol) = 45.980 g

$2\text{H}_2\text{O}$: (4 mol)(1.008 g/mol) +
(2 mol)(15.999 g/mol) = 36.030 g

2NaOH : (2 mol)(22.990 g/mol) +
(2 mol)(15.999 g/mol) + (2 mol)(1.008 g/mol)
= 79.994 g

H_2 : (2 mol)(1.008 g/mol) = 2.016 g

45.980 g Na + 36.030 g $\text{H}_2\text{O} \rightarrow$

79.994 g NaOH + 2.016 g H_2

82.01 g reactants = 82.01 g products

Chemistry Study Guide Answers Stoichiometry

Dixie Goss



Chemistry Study Guide Answers Stoichiometry:

Study Guide/Selected Solutions Manual Julie R. Frentrup, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

Basic Concepts of Chemistry, Study Guide Leo J. Malone, 2000 The text s three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems

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remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students Resources in Education ,1980 Study Guide for Chemistry by Steven S. Zumdahl Martha B. Barrett,1986

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