

Chemistry 105 Practice Final Exam KEY

1. A pure substance
 - a. Must be composed of atoms of the same type.
 - b. Cannot be separated into simpler substances by physical means.
 - c. Must be a compound.
 - d. Has different chemical properties depending on its source.
 - e. Can have a composition that varies from sample to sample.
2. Which of the following represents a physical change?
 - a. Milk turns sour.
 - b. Rust forms on iron nails.
 - c. Sugar ferments to form ethanol.
 - d. An egg begins to smell very bad.
 - e. Sugar melts and forms a syrupy liquid.
3. The densities of ethylene glycol and of mercury are 1.11 and 13.5 g/mL, respectively. What volume of mercury has the same mass as 125 mL ethylene glycol?
 - a. 10.3 mL
 - b. 8.34 mL
 - c. 0.120 mL
 - d. 0.0971 mL
 - e. 1870 mL
4. Gallium has two naturally occurring isotopes with the following masses and natural abundances. Calculate the average atomic mass of Ga.

^{69}Ga	68.9256 amu	60.108%
^{71}Ga	70.9247 amu	39.892%

 - a. 69.925 amu
 - b. 70.127 amu
 - c. 70.000 amu
 - d. 69.824 amu
 - e. 69.723 amu
5. Which transition in a hydrogen atom involves the *largest* change in energy?
 - a. $n=2$ to $n=3$
 - b. $n=3$ to $n=4$
 - c. $n=1$ to $n=3$
 - d. $n=1$ to $n=2$
 - e. $n=9$ to $n=10$
6. The wavelength of a photon of light varies:
 - a. Inversely as the frequency and inversely as the energy
 - b. Inversely as the frequency and directly as the energy
 - c. Directly as the frequency and inversely as the energy

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