

A. Covalent & Ionic Compounds/Bonds

1. When an atom loses an electron it becomes a positive ion (cation), and when it gains an electron it becomes a negative ion (anion).

2. An element in the halogen group has 7 valence electrons, and wants to gain 1 more to have 8. This is because of the octet rule.

3. An ionic bond is the attraction between oppositely charged ions, AKA a metal and a nonmetal.

4. In a covalent bond, atoms share electrons, whereas in an ionic bond, atoms transfer electrons.

5. (Honors only) When electrons are shared unequally in a covalent bond it is polar, and when they're shared equally it is nonpolar. You can predict if a compound is ionic, polar covalent or nonpolar covalent by subtracting the electronegativities.

6. An atom can have up to 8 valence electrons. The family (column) on the periodic table that all have 8 valence electrons is called noble gases.

7. Atoms can gain, lose, or share valence electrons to become more stable.

B. Formulas for Ionic compounds - Fill in the chart below

	Magnesium	Fluorine	Nitrogen	Lithium
# OF VALENCE ELECTRONS	<u>2</u>	<u>7</u>	<u>5</u>	<u>1</u>
OXIDATION #	<u>+2</u>	<u>-1</u>	<u>-3</u>	<u>+1</u>

2. What is the **chemical formula** for each of the following ionic compounds?

(Hint: you can either draw the Lewis structures and use arrows to show the transfer of electrons, OR use the periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.)

- Potassium and Chlorine - KCl
- Calcium and Bromine - CaBr₂
- Magnesium and oxygen - MgO

Chemical Bonds Study Guide

**John Krenos, Joseph Potenza, Carl
Hoeger**

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