

CHEM 110
Extra In-Class Practice Problems
Lecture Set 3

1. Answer the following questions dealing with Avogadro's number.

A. How many molecules present in 2.0 mol CH₄?

$$\frac{2.0 \text{ mol CH}_4}{1 \text{ mol CH}_4} \left| \frac{6.022 \times 10^{23} \text{ molecules}}{1 \text{ molecule}} \right. = \boxed{1.2 \times 10^{24} \text{ molecules CH}_4}$$

How many atoms are present?

$$\frac{1.2 \times 10^{24} \text{ molecules CH}_4}{1 \text{ molecule CH}_4} \left| \frac{5 \text{ atoms}}{1 \text{ molecule CH}_4} \right. = \boxed{6.0 \times 10^{24} \text{ atoms}}$$

B. How many formula units of CaCl₂ are present in 0.543 mol CaCl₂?

$$\frac{0.543 \text{ mol CaCl}_2}{1 \text{ mol CaCl}_2} \left| \frac{6.022 \times 10^{23} \text{ f.u.}}{1 \text{ molecule}} \right. = \boxed{3.27 \times 10^{23} \text{ f.u. CaCl}_2}$$

C. 9.81×10^{21} molecules O₂ is how many moles?

$$\frac{9.81 \times 10^{21} \text{ molecules O}_2}{1 \text{ molecule}} \left| \frac{1 \text{ mol O}_2}{6.022 \times 10^{23} \text{ molecules}} \right. = \boxed{0.0163 \text{ mol O}_2}$$

D. How many atoms are present in 0.371 mol P₂O₅?

$$\frac{0.371 \text{ mol P}_2\text{O}_5}{1 \text{ mol P}_2\text{O}_5} \left| \frac{6.022 \times 10^{23} \text{ molecules P}_2\text{O}_5}{1 \text{ molecule}} \right. \left| \frac{7 \text{ atoms}}{1 \text{ molecule P}_2\text{O}_5} \right. = \boxed{1.56 \times 10^{24} \text{ atoms}}$$

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JS Bruner

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