

Chemistry 105 Equation Sheet

Constants

$$\begin{aligned}
 a_0 &= 0.5291772059 \text{ \AA} \\
 c &= 2.99792458 \times 10^8 \text{ m/s} \\
 e &= 1.6021765 \times 10^{-19} \text{ C} \\
 \epsilon_0 &= 8.8541877 \times 10^{-12} \text{ F/m} \\
 \hbar &= 6.626 \times 10^{-34} \text{ J}\cdot\text{s} \\
 k_B &= 1.38065 \times 10^{-23} \text{ J/K} \\
 k_e &= 9.987551 \times 10^9 \frac{\text{N}\cdot\text{m}^2}{\text{C}^2} \\
 m_e &= 9.109393 \times 10^{-31} \text{ kg} \\
 m_p &= 1.6749273 \times 10^{-27} \text{ kg} \\
 m_\pi &= 1.6726217 \times 10^{-27} \text{ kg} \\
 N_A &= 6.022142 \times 10^{23} \text{ mol}^{-1} \\
 R &= 8.31447 \frac{\text{J}}{\text{mol}\cdot\text{K}} \\
 &= 0.0820575 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}} \\
 &= 1.987207 \frac{\text{cal}}{\text{mol}\cdot\text{K}} \\
 R_N &= 1.0974 \times 10^7 \text{ m}^2
 \end{aligned}$$

Conversions (* = exact)

$$\begin{aligned}
 \text{Length} \\
 1 \text{ in} &= 2.54 \text{ cm}^* \\
 1 \text{ ft} &= 12 \text{ inches}^* \\
 1 \text{ yd} &= 3 \text{ ft}^* \\
 1 \text{ m} &= 39.37 \text{ inches}^* \\
 1 \text{ mile} &= 5280 \text{ feet}^* \\
 1 \text{ mile} &= 1.609 \text{ km} \\
 \text{Mass} \\
 1 \text{ lb} &= 453.59237 \text{ g} \\
 &= 16 \text{ oz avdp.}^* \\
 1 \text{ US lb} &= 12 \text{ troy ounces}^* \\
 1 \text{ ton} &= 2000 \text{ lb}^* \\
 1 \text{ amu} &= 1.660538 \times 10^{-24} \text{ g}
 \end{aligned}$$

Volume

$$\begin{aligned}
 1 \text{ US gal} &= 3.7854 \text{ L} = 4 \text{ US qt}^* \\
 1 \text{ US qt} &= 32 \text{ US fl oz.}^* \\
 1 \text{ US fluid oz.} &= 29.5735 \text{ mL} \\
 1 \text{ cm}^3 &= 0.001 \text{ L}^* \\
 1 \text{ in}^3 &= 16.4 \text{ cm}^3
 \end{aligned}$$

* = exact ν = frequency v = velocity

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Pressure

$$\begin{aligned}
 1 \text{ atm} &= 760 \text{ torr}^* \\
 &= 101325 \text{ Pa}^* \\
 &= 406.733 \text{ in H}_2\text{O} \\
 &= 14.6959 \text{ lb/in}^2 \text{ (psi)} \\
 1 \text{ torr} &= 1 \text{ mm Hg (at } 0^\circ\text{C})^* \\
 1 \text{ Pa} &= 1 \text{ kg}\cdot\text{m}^{-1}\cdot\text{s}^{-2} \\
 1 \text{ bar} &= 100,000 \text{ Pa}^*
 \end{aligned}$$

Force

$$1 \text{ N} = 1 \text{ kg}\cdot\text{m}\cdot\text{s}^{-2}$$

Power

$$1 \text{ Watt} = 1 \text{ J/s}$$

Energy

$$\begin{aligned}
 1 \text{ J} &= 1 \text{ kg}\cdot\text{m}^2/\text{s}^2 = 0.2390 \text{ cal} \\
 1 \text{ J} &= 1 \text{ V}\cdot\text{C} = 0.009869 \text{ L}\cdot\text{atm} \\
 1 \text{ eV} &= 1.602 \times 10^{-19} \text{ J} \\
 1 \text{ L}\cdot\text{atm} &= 101.3 \text{ J}
 \end{aligned}$$

Formulas

$$\rho = m/V$$

$$T_F = \frac{9F}{5C} T_C + 32 \text{ }^\circ\text{F}$$

$$T_C = \frac{5F}{9C} T_F + 273.15 \text{ K}$$

$$c = \lambda \nu$$

$$E = h\nu$$

$$\lambda = \frac{h}{mc}$$

$$E_{\text{kinetic}} = \frac{1}{2}mv^2$$

$$E = mc^2$$

$$\Phi = h\nu - KE_{\text{electron}}$$

$$E_n = -Z^2 \frac{R_N hc}{n^2}$$

$$\Delta x \Delta(mv) \geq \frac{\hbar}{4\pi}$$

$$\frac{1}{\lambda} = R_N \left(\frac{1}{n_1^2} - \frac{1}{n_2^2} \right)$$

$$V(r) = \frac{Q_1 Q_2}{4\pi \epsilon_0 r^2}$$

$$F(r) = k_s \frac{Q_1 Q_2}{r^2}$$

$$\mu = Qr$$

$$\Delta E = q + w$$

$$\Delta E_{\text{int}} = 0$$

$$\Delta S_{\text{int}} \geq 0$$

$$S = k_B \ln W$$

$$q = mC\Delta T$$

$$w = -P_{\text{ext}} \Delta V$$

$$\Delta H = \Delta E + P\Delta V$$

$$\Delta S = \frac{Q_{\text{rev}}}{T}$$

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta G = \Delta G^\circ + RT \ln Q$$

$$\Delta H_{\text{rxn}} = \Sigma n H_f^{\circ,\text{prod}} - \Sigma m H_f^{\circ,\text{react}}$$

$$\Delta H_{\text{rxn}} \approx \Sigma n D_{\text{broken}} - \Sigma m D_{\text{formed}}$$

$$PV = nRT$$

$$P_{\text{tot}} = P_1 + P_2 + P_3 + \dots$$

$$\frac{\text{Rate}_A}{\text{Rate}_B} = \frac{\text{Time}_B}{\text{Time}_A} = \sqrt{\frac{M_B}{M_A}}$$

$$KE_{\text{ave}} = \frac{3}{2}RT$$

$$v_{\text{rms}} = \sqrt{v^2} = \sqrt{\frac{3RT}{M}}$$

$$\text{pX} = -\log X$$

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-]$$

Chemistry Equation Sheet

Mary Millhollen, Richard H. Langley

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101 Ways to Score Higher on Your MCAT Marti Anne Maguire,Paula Stiles,2010

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Combustion Theory Forman A. Williams,2018-03-05

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Table of Contents Chemistry Equation Sheet

1. Understanding the eBook Chemistry Equation Sheet
 - The Rise of Digital Reading Chemistry Equation Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Equation Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Equation Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Equation Sheet
 - Personalized Recommendations
 - Chemistry Equation Sheet User Reviews and Ratings
 - Chemistry Equation Sheet and Bestseller Lists
5. Accessing Chemistry Equation Sheet Free and Paid eBooks

- Chemistry Equation Sheet Public Domain eBooks
- Chemistry Equation Sheet eBook Subscription Services
- Chemistry Equation Sheet Budget-Friendly Options

6. Navigating Chemistry Equation Sheet eBook Formats

- ePUB, PDF, MOBI, and More
- Chemistry Equation Sheet Compatibility with Devices
- Chemistry Equation Sheet Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Chemistry Equation Sheet
- Highlighting and Note-Taking Chemistry Equation Sheet
- Interactive Elements Chemistry Equation Sheet

8. Staying Engaged with Chemistry Equation Sheet

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Equation Sheet

9. Balancing eBooks and Physical Books Chemistry Equation Sheet

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Chemistry Equation Sheet

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Chemistry Equation Sheet

- Setting Reading Goals Chemistry Equation Sheet
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemistry Equation Sheet

- Fact-Checking eBook Content of Chemistry Equation Sheet
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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