

### Good Friday Worksheet

Period 4.8.8: Classification (Homework, 10 marks)

1. Identify which type of bond has the following properties.  
 (Answer: Ionic, Metallic, Covalent)

Metallic bond	Ionic bond	Covalent bond
<ul style="list-style-type: none"> <li>Use of delocalized electrons</li> <li>Conduction when solid</li> </ul>	<ul style="list-style-type: none"> <li>Oppositive charges</li> <li>Large resulting ionic radius</li> <li>Electrostatic attraction between delocalized electrons</li> </ul>	<ul style="list-style-type: none"> <li>Clear ion solid, liquids, or gases</li> <li>Smaller resulting ionic radius</li> <li>Stronger intermolecular interactions</li> </ul>

2. Circle the correct word that will complete the sentence below.

Metallic bonds result from ~~electron~~ electro-negativity differences. The element with a stronger electro-negativity will ~~group~~ electrons from the element with a weaker electro-negativity. Covalent bonds result in ~~group~~ electro-negativity differences. So, the elements will ~~group~~ their electrons with each other to fulfill their ~~group~~ requirements.

3. Fill in the chart below using the element names or compound names.

Symbol or Compound	Formula	Name	Lewis Structure	ip Covalency	ip Covalency Change
Tin(II)	$\text{Sn}_{2+}$	Stannous Iodide	$\text{[SnI}_3]^{2-}$	Nonpolar	Nonpolar
Covalent	$\text{C}_6\text{H}_6$	Carbon hexafluoride	$\text{[C]F}_6$	Polar	Tetrahedral
Tin(IV)	$\text{SnO}_2$	Copper(II) iodide	$\text{[SnI}_3]^{2-}$	Polar	Polar
Covalent	$\text{O}_2$	Oxygen difluoride	$\text{[O]F}_2$	Polar	Linear
Tin(II)	$\text{Sn}_{2+}$	Stannous hexafluoride	$\text{[SnF}_6]^{2-}$	Nonpolar	Nonpolar
Tin(IV)	$\text{SnO}_2$	Copper(II) iodide	$\text{[SnI}_3]^{2-}$	Polar	Polar
Covalent	$\text{O}_2$	Oxygen trifluoride	$\text{[O]F}_3$	Polar	Trigonal planar
Covalent	$\text{H}_2\text{S}$	Hydrogen bromide	$\text{H} - \text{S} - \text{H}$	Polar	Linear

### Electro-Static

#### Important Vocabulary & their Definitions

- Electrostatic: ~~long~~ ~~at~~ ~~one~~ ~~other~~ ~~one~~ ~~other~~ ~~one~~ ~~other~~ ~~one~~ ~~other~~ ~~one~~ ~~other~~
- Proton - ~~group~~ ~~at~~ ~~one~~ ~~other~~

1. Circle the correct word that will complete the sentence below.

Elements with a large ~~convention~~ will have a ~~large~~ frequency and ~~large~~ energy.

2. Calculate the energy (E) and wavelength (λ) of a photon of light with a frequency (f) of  $5.165 \times 10^{14} \text{ Hz}$ .

$$E = hf$$

$$= (6.626 \times 10^{-34} \text{ J})(5.165 \times 10^{14} \text{ Hz})$$

$$E = 3.39 \times 10^{-19} \text{ J} = \frac{3.39}{1.602 \times 10^{-19} \text{ J}}$$

$$\lambda = \frac{c}{f} = \frac{3.0 \times 10^8 \text{ m/s}}{5.165 \times 10^{14} \text{ Hz}} = 5.8 \times 10^{-7} \text{ m}$$

$$E = 6.626 \times 10^{-34} \text{ J} \times 5.165 \times 10^{14} \text{ Hz}$$

$$E = 3.39 \times 10^{-19} \text{ J}$$

# Chemistry Final Study Guide Answer Key

**Thomas Elliott Taylor**

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