

Chemistry Study Guide: Unit 6: Gases

Kinetic Molecular Theory (KMT): a model or theory that is used to explain the behavior of gases.

- A. Gases contain particles (molecules or atoms) that are in constant, random, straight-line motion.
- B. Gas particles collide with each other and with the walls of the container. These collisions are perfectly elastic, meaning kinetic energy (energy an object has because of its motion) is being transferred among the particles but no kinetic energy is lost.
- C. Gas particles are separated by relatively great distances. Because of this, the volume of the particles themselves is negligible (insignificant).
- D. Gas particles do not attract or repel each other. This is because they are far away from each other.

Gas Pressure: Force exerted by a gas on the walls of its container.

- Caused by collisions between gas particles and walls of the container.
- Units: kilopascal (kPa), Atmosphere (atm), millimeters mercury (mm Hg)
- 1 atm = 760mmHg = 101.3kPa
- Found on table A with temperature on periodic table

Temperature: a measure of the average kinetic energy of the particles in a substance.

- As average kinetic energy of particles in a substance increases, temperature of a substance increases.
- In science, Celsius and kelvin are common temperature scales.
- $K = C^\circ + 273$, $C^\circ = K - 273$
- Temperature is not a form of energy.

Relationship between pressure and volume of a gas:

Boyle's Law: Pressure and Volume (Temperature Constant)

If the temperature is constant, then the pressure of a gas increases, as the volume decreases. As the pressure decreases the volume increases.

- Inverse relationship between pressure and volume.

Equation: $P_1 \cdot V_1 = P_2 \cdot V_2$

Charles Law: Temperature and Volume (Pressure Constant)

If the pressure is constant, then as the temperature of an enclosed gas increases, the volume increases. As the temperature decreases, the volume decreases.

- Direct relationship between temperature and volume
- Temperature must be converted to Kelvin.

Equation: $V_1 / T_1 = V_2 / T_2$

Gay-Lussac's Law: Temperature and Pressure (Volume is constant)

If the volume is constant, then as the temperature of an enclosed gas increases, the pressure increases. As the temperature decreases, the pressure decreases.

- Direct relationship between temperature and pressure.
- Temperature must be converted to Kelvin.

Equation: $P_1 / T_1 = P_2 / T_2$

Chemistry Gases Unit Study Guide

Darrell D. Ebbing, Susan M. Schelble

Chemistry Gases Unit Study Guide:

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