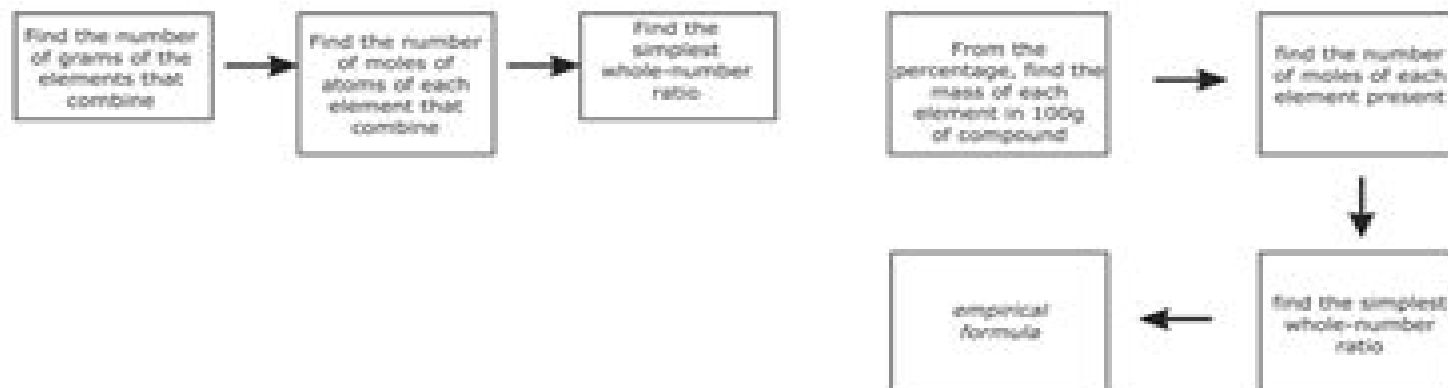


RATIOS ...	Carbon ( $A_r = 12$ )	Hydrogen ( $A_r = 1$ )	Chlorine ( $A_r = 35.5$ )	Comments and tips
Reacting mass or % mass	24.24	4.04	71.72	just think of it as based on 100g
molar ratio (mass in g / $A_r$ )	$24.24 / 12 = 2.02 \text{ mol}$	$4.04 / 1 = 4.04 \text{ mol}$	$71.72 / 35.5 = 2.02 \text{ mol}$	can now divide by smallest ratio number
simplest whole number ratio	$2.02 / 2.02 = 1$	$4.04 / 2.02 = 2$	$2.02 / 2.02 = 1$	
therefore the simplest ratio = empirical formula for the chlorinated hydrocarbon = $\text{CH}_2\text{Cl}$				
BUT the molecular mass is 99, and the empirical formula mass is 49.5 ( $12+2+35.5$ )				
AND $99 / 49.5 = 2$ , and so the molecular formula must be $2 \times \text{CH}_2\text{Cl} = \text{C}_2\text{H}_4\text{Cl}_2$				



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**CH Cherryholmes**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'CH Cherryholmes' text.

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