

Worksheet 1.6

Supplemental Instruction
Iowa State University

Leader: Lea

Course: Chem 177

Instructor: Dr. Schewe-Miller

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- 1.) A piece of aluminum foil 1.00 cm square and 0.550 mm thick is allowed to react with bromine to form aluminum bromide.

a.) How many moles of aluminum were used? (density Al = 2.699 g/cm³)

$$.550 \text{ mm} \times \frac{1 \text{ cm}}{10 \text{ mm}} = .055 \text{ cm}$$

$$1 \text{ cm} \times .055 \text{ cm} = .055 \text{ cm}^2 \times \frac{2.699 \text{ g}}{\text{cm}^3} = .148445 \text{ g Al} \times \frac{1 \text{ mol Al}}{26.98 \text{ g Al}} = .005502 \text{ mol Al}$$

$5.50 \times 10^{-3} \text{ mol Al}$

b.) How many grams of aluminum bromide form, assuming the aluminum reacts completely? What type of reaction is this?



$$5.50 \times 10^{-3} \text{ mol Al} \times \frac{2 \text{ mol AlBr}_3}{2 \text{ mol Al}} \times \frac{266.69 \text{ g AlBr}_3}{1 \text{ mol AlBr}_3} = 1.47 \text{ g}$$

composition reaction

2.) The chemical formula for benzene is C₆H₆.

a. Determine the percent composition of H and C for benzene.

$$6(12.01) + 6(1.01) = 78.12$$

$$\text{C}: \frac{72.06}{78.12} \times 100 = 92\% \quad \text{H}: 100 - 92\% = 8\%$$

$\text{C}: 92\% \\ \text{H}: 8\%$

b. Write a balanced combustion reaction for benzene.



c. Assuming you have 364g of benzene, how many moles of O₂ will be required to completely combust this amount of benzene?

$$364 \text{ g} \times \frac{1 \text{ mol C}_6\text{H}_6}{78.12 \text{ g C}_6\text{H}_6} \times \frac{15 \text{ mol O}_2}{2 \text{ mol C}_6\text{H}_6} = 35.0 \text{ mol}$$

d. How many grams of CO₂ will this produce? How many grams of H₂O?

$$\text{CO}_2: 364 \text{ g C}_6\text{H}_6 \times \frac{1 \text{ mol C}_6\text{H}_6}{78.12 \text{ g C}_6\text{H}_6} \times \frac{12 \text{ mol CO}_2}{2 \text{ mol C}_6\text{H}_6} \times \frac{44.01 \text{ g CO}_2}{1 \text{ mol CO}_2} = 1.23 \times 10^3 \text{ g CO}_2$$

$$\text{H}_2\text{O}: 364 \text{ g C}_6\text{H}_6 \times \frac{1 \text{ mol C}_6\text{H}_6}{78.12 \text{ g C}_6\text{H}_6} \times \frac{6 \text{ mol H}_2\text{O}}{2 \text{ mol C}_6\text{H}_6} \times \frac{18 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 2.52 \times 10^2 \text{ g H}_2\text{O}$$

e. You perform the combustion, and 956g of CO₂ are produced. What is the percent yield?

$$\frac{\text{Actual}}{\text{Theoretical}} \times 100 = \frac{956 \text{ g CO}_2}{1230 \text{ g CO}_2} \times 100 = 77.7\%$$

Chemistry Stoichiometry Problem Sheet 2 Answer Key

Jin-Ying Zhang



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1. Understanding the eBook Chemistry Stoichiometry Problem Sheet 2 Answer Key
 - The Rise of Digital Reading Chemistry Stoichiometry Problem Sheet 2 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Stoichiometry Problem Sheet 2 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Stoichiometry Problem Sheet 2 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Stoichiometry Problem Sheet 2 Answer Key

- Personalized Recommendations
 - Chemistry Stoichiometry Problem Sheet 2 Answer Key User Reviews and Ratings
 - Chemistry Stoichiometry Problem Sheet 2 Answer Key and Bestseller Lists
5. Accessing Chemistry Stoichiometry Problem Sheet 2 Answer Key Free and Paid eBooks
 - Chemistry Stoichiometry Problem Sheet 2 Answer Key Public Domain eBooks
 - Chemistry Stoichiometry Problem Sheet 2 Answer Key eBook Subscription Services
 - Chemistry Stoichiometry Problem Sheet 2 Answer Key Budget-Friendly Options
 6. Navigating Chemistry Stoichiometry Problem Sheet 2 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Stoichiometry Problem Sheet 2 Answer Key Compatibility with Devices
 - Chemistry Stoichiometry Problem Sheet 2 Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Stoichiometry Problem Sheet 2 Answer Key
 - Highlighting and Note-Taking Chemistry Stoichiometry Problem Sheet 2 Answer Key
 - Interactive Elements Chemistry Stoichiometry Problem Sheet 2 Answer Key
 8. Staying Engaged with Chemistry Stoichiometry Problem Sheet 2 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Stoichiometry Problem Sheet 2 Answer Key
 9. Balancing eBooks and Physical Books Chemistry Stoichiometry Problem Sheet 2 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Stoichiometry Problem Sheet 2 Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Stoichiometry Problem Sheet 2 Answer Key
 - Setting Reading Goals Chemistry Stoichiometry Problem Sheet 2 Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Stoichiometry Problem Sheet 2 Answer Key

- Fact-Checking eBook Content of Chemistry Stoichiometry Problem Sheet 2 Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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