

## Review for Ch. 8 Test

**8.1 Describing Chemical Reactions**

- know how to write word equations (i.e. you have to be able to name compounds, including acids, ionic, and covalent compounds from Ch. 7)
- know how to write a formula equation
- What 2 things do coefficients represent in an equation? **Relative number of moles of a substance ; relative number of molecules needed for a reaction to occur**
- know common symbols used in a reaction
- know 4 evidences of a chemical reaction:
  1. **energy in the form of heat or light**
  2. **production of a gas**
  3. **change in color**
  4. **formation of a precipitate**
- know the **Law of conservation of matter; matter can neither be created nor destroyed in a chemical reaction (which means, mass of reactants must equal the mass of products)**
- know what a chemical equation represents: **relative amounts of reactants and products**
- know how to balance chemical equations (combustion reactions are particularly tricky, practice!)

1. Complete the equation and write a word equation for the following reaction:  $\text{ZnO} + \text{FeSO}_4 \rightleftharpoons$   
 $\text{ZnO} + \text{FeSO}_4 \rightleftharpoons \text{ZnSO}_4 + \text{FeO}$

**Zinc oxide reacts with iron (II) sulfate to yield zinc sulfate and iron (II) oxide.**

2. Write a formula equation for the reaction of a solution of  $\text{AgCl}$  with potassium metal:



3. What elements are always diatomic gases at room temperature?

**All of the halogens (Group 17) ,  $\text{H}_2$ ,  $\text{N}_2$ ,  $\text{O}_2$**

4. Can you balance these reactions:



# Describing Chemical Reactions Study Guide Answer Key

**Paul B. Kelter, Steven S. Zumdahl**

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