

## Chemistry — Unit 8 Test

### 1. The following graph shows the relationship between the number of atoms and the mass of an element.

- a. The two axes for this graph are the
- b. This graph shows a linear relationship between the number of atoms and the mass of an element.
- c. This graph shows a non-linear relationship between the number of atoms and the mass of an element.
- d. This graph shows a non-linear relationship between the number of atoms and the mass of an element.
- e. This graph shows a non-linear relationship between the number of atoms and the mass of an element.

### 2. The following diagrams show three different ways to represent the same element.

Which of the following best shows the correct relationship between the three representations for this element?

- a. This diagram shows the correct relationship between the three representations for this element.
- b. This diagram shows the correct relationship between the three representations for this element.
- c. This diagram shows the correct relationship between the three representations for this element.
- d. This diagram shows the correct relationship between the three representations for this element.
- e. This diagram shows the correct relationship between the three representations for this element.



### 3. Relative reactivities of metals are often determined by

- a. the time that each metal takes to start reacting with a dilute acid such as 1M  $\text{HCl}$ . The greater the time, the higher the metal's reactivity.
- b. the time that each metal takes to start reacting with a dilute acid such as 1M  $\text{HCl}$ . The greater the time, the lower the metal's reactivity.
- c. the time that each metal takes to start reacting with a dilute acid such as 1M  $\text{HCl}$ . The greater the time, the higher the metal's reactivity.
- d. the time that each metal takes to start reacting with a dilute acid such as 1M  $\text{HCl}$ . The greater the time, the lower the metal's reactivity.
- e. the time that each metal takes to start reacting with a dilute acid such as 1M  $\text{HCl}$ . The greater the time, the higher the metal's reactivity.

The reactivities of metals can also be determined by their rates of displacement reactions with  $\text{H}_2\text{O}$ .

- a. Potassium reacts with  $\text{H}_2\text{O}$  to form  $\text{K}_2\text{O}$  and  $\text{H}_2$ . This is an example of a displacement reaction.
- b. Potassium reacts with  $\text{H}_2\text{O}$  to form  $\text{K}_2\text{O}$  and  $\text{H}_2$ . This is an example of a double displacement reaction.
- c. Potassium reacts with  $\text{H}_2\text{O}$  to form  $\text{K}_2\text{O}$  and  $\text{H}_2$ . This is an example of a double displacement reaction.
- d. Potassium reacts with  $\text{H}_2\text{O}$  to form  $\text{K}_2\text{O}$  and  $\text{H}_2$ . This is an example of a double displacement reaction.
- e. Potassium reacts with  $\text{H}_2\text{O}$  to form  $\text{K}_2\text{O}$  and  $\text{H}_2$ . This is an example of a double displacement reaction.

# Chemistry Unit 9 Test

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