

# Moles and Reactions

Student

Date

Page

## Introduction: Molar Ratios in Stoichiometry

Stoichiometric quantities (moles) are calculated from the mass (grams) of the substances involved in the reaction.



The molar ratio of the reactants and products in the reaction is determined by the stoichiometric coefficients. In the above reaction, the molar ratio of the reactants is 2:1 and the molar ratio of the products is 2:1. This means that for every 2 moles of  $\text{H}_2$  that react, 1 mole of  $\text{O}_2$  reacts to produce 2 moles of  $\text{H}_2\text{O}$ . Similarly, the molar ratio of the products is 2:1, which means that for every 2 moles of  $\text{H}_2\text{O}$  that are produced, 1 mole of  $\text{O}_2$  is consumed. The molar ratio of the reactants and products in a chemical reaction are related by the ratio of the coefficients of the substances.

## Critical Thinking Questions

Using the equations 1-4, write balanced chemical equations for the reactions given.

1. a) Write a balanced equation for the reaction of 1.00 mole of  $\text{H}_2$  with 1.00 mole of  $\text{O}_2$  to produce 2.00 moles of  $\text{H}_2\text{O}$ .



- b) Write a balanced equation for the reaction of 1.00 mole of  $\text{H}_2$  with 1.00 mole of  $\text{O}_2$  to produce 1.00 mole of  $\text{H}_2\text{O}$ .



2. Write a balanced equation for the reaction of 1.00 mole of  $\text{H}_2$  with 1.00 mole of  $\text{O}_2$  to produce 1.00 mole of  $\text{H}_2\text{O}$ .



3. Write a balanced equation for the reaction of 1.00 mole of  $\text{H}_2$  with 1.00 mole of  $\text{O}_2$  to produce 1.00 mole of  $\text{H}_2\text{O}$ .

The balanced equation is  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ . 1 mole of  $\text{H}_2$  reacts with 0.5 mole of  $\text{O}_2$  to produce 1 mole of  $\text{H}_2\text{O}$ .

4. Write a balanced equation for the reaction of 1.00 mole of  $\text{H}_2$  with 1.00 mole of  $\text{O}_2$  to produce 1.00 mole of  $\text{H}_2\text{O}$ .



5. Write a balanced equation for the reaction of 1.00 mole of  $\text{H}_2$  with 1.00 mole of  $\text{O}_2$  to produce 1.00 mole of  $\text{H}_2\text{O}$ .



# Chemquest 42 Answers

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