

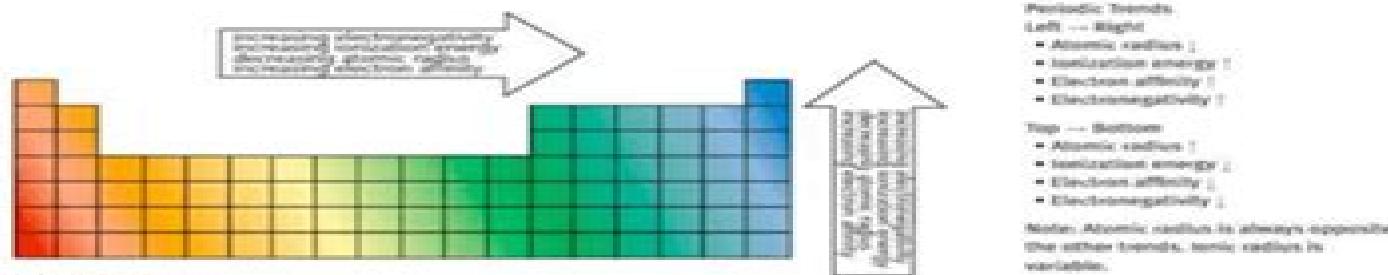
Chemistry:

CHAPTER 1

- Charge of electron $\approx 1.6 \times 10^{-19}$; Avogadro's # $\approx 6 \times 10^{23}$
- Mass number (A) on top of element, atomic number (Z) below
- $E = \frac{hc}{\lambda} = -R \left[\frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$; $E = qV$ (energy = charge on e- x electron volts); $c = \lambda f$ ($c = 3 \times 10^8$)
- Heisenberg uncertainty — impossible to determine momentum and position simultaneously
- Pauli exclusion — no electrons have same quantum #’s
- Max e- in a shell $\approx 2n^2$, max e- in subshell $\approx 4l+2$
 - Possible values for l ≈ 0 to $n-1$, possible values for ml $\approx -l$ to $+l$
- S can hold 2 e-, p ≈ 6 , d ≈ 10 , f ≈ 14
- Aufbau principle (building up) — lowest n+l will fill first
- For anions add extra e-, for cations remove e- from subshells with highest n first
- Hund’s — half filled orbitals with parallel spin e-, attracted to magnetic field
 - Diamagnetic — paired e-, repelled by magnetic field
 - Ferromagnetic — unpaired e-, randomly oriented dipoles, become magnetized in field
- A faraday is equal to one mole of electric charge

CHAPTER 2

- Z_{eff} increases from left to right, constant up and down
- Ionic radius is smallest near metalloid line and increases out
- F is smallest, most EN, highest IE and EA; Cs is the largest and least EN
- Complex ions from transition metals associate with water to form **hydration complexes**
- Atomic radius has the opposite trend of all others



CHAPTER 3

- C, N, O, F, Na, Mg always abide by octet rule; H, He, Li, Be, B can have incomplete; period 3 or higher can have expanded
- **Coordinate covalent bond** — both e- are contributed by one atom (lone pair attack, acid-base)
- Ionic ΔEN difference greater than 1.7, polar covalent ≈ 0.5 to 1.7, nonpolar ΔEN less than 0.5
 - Ionic are stronger than covalent (which are both stronger than IM forces)
 - Van der Waals occurs more with close proximity and large molecules
- **Dipole moment** $\approx p = qd$
- Triple bonds are shorter and have more energy
- Major resonance contributor is more stable (less formal charge, negative on more EN atom)
 - Formal charge = valence - nonbonding - $\frac{1}{2}$ bonding
- Electron geometry is when bonding and lone are equal, molecular geo is when lone repel more

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Answers are included in the Appendix **Understanding Earth Student Study Guide** Peter L. Kresan, Reed Mencke, 2006-05-03 The guide helps students prepare for lectures and exams with a heavy emphasis on utilizing the book's Web resources **Study Guide and Full Solutions Manual** Susan McMurry, John McMurry, 2002-08 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all end of chapter problems in the textbook **Students Study Guide** Fredene, Pearson, 2003-07 **Chemistry Equations & Answers** Speedy Publishing, 2014-08-26 General chemistry inorganic chemistry organic chemistry and biochemistry are all difficult courses requiring much memorization for the student Essentially there is no easy way to learn formulas and facts This is why chemistry classes are such challenges to students even the best ones However a chemistry equations and answers study guide can help the student When used as a quick reference guide it can be used often to determine the formulas needed for various questions The astute student can cleverly devise ways to make the guide useful for test questions or other circumstances requiring one of the many chemistry equations

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