

GENERAL CHEMISTRY REFERENCE SHEET

Common Metric Prefixes and Units		
Unit	Symbol	Conversions
meter	m	1 m = 39.37 inches
centimeter	cm	1 inch = 2.54 cm exactly
Angstrom	Å	1 Å = 1×10^{-10} m
gram	g	1 g = 0.002205 lb
liter	L	1 L = 1000 cm ³ = 1.056710 qt
Kelvin	K	K = °C + 273.15
kilo	k	10 ³
deci	d	10 ⁻¹
centi	c	10 ⁻²
milli	m	10 ⁻³
micro	μ	10 ⁻⁶
nano	n	10 ⁻⁹
pico	p	10 ⁻¹²

Nomenclature

Common Cations		
Charge	Formula	Name
1+	H ⁺	hydrogen ion
	H ₃ O ⁺	hydronium ion
	Alkali metal ⁺	alkali metal ion
	NH ₄ ⁺	ammonium ion
	Ag ⁺	silver ion
2+	Cu ⁺	copper(I)
	Alkaline earth metal ²⁺	alkaline earth metal ion
	Zn ²⁺	zinc ion
	Co ²⁺	cobalt(II)
	Cu ²⁺	copper(II)
	Fe ²⁺	iron(II)
	Mn ²⁺	manganese(II)
	Hg ₂ ²⁺	mercury(I)
	Hg ₂ ²⁺	mercury(II)
	Ni ²⁺	nickel(II)
	Pb ²⁺	lead(II)
	Sn ²⁺	tin(II)
3+	Al ³⁺	aluminum ion
	Cr ³⁺	chromium(III)
	Fe ³⁺	iron(III)

Common Anions		
Charge	Formula	Name
1-	H ⁻	hydride ion
	F ⁻	fluoride ion
	Cl ⁻	chloride ion
	Br ⁻	bromide ion
	I ⁻	iodide ion
	CN ⁻	cyanide ion
	OH ⁻	hydroxide ion
	NO ₃ ⁻	nitrate
	NO ₂ ⁻	nitrite
	HCO ₃ ⁻	hydrogen carbonate or bicarbonate ion
	HSO ₄ ⁻	hydrogen sulfate ion

Charge	Formula	Name
1-	ClO ₄ ⁻	perchlorate ion
	ClO ₃ ⁻	chlorate ion
	ClO ₂ ⁻	chlorite ion
	ClO ⁻	hypochlorite ion
	MnO ₄ ⁻	permanganate ion
	H ₂ PO ₄ ⁻	dihydrogen phosphate ion
	C ₂ H ₃ O ₂ ⁻ (or CH ₃ COO ⁻)	acetate ion
2-	O ²⁻	oxide ion
	S ²⁻	sulfide ion
	SO ₄ ²⁻	sulfate ion
	SO ₃ ²⁻	sulfite ion
	CO ₃ ²⁻	carbonate ion
	Cr ₂ O ₇ ²⁻	dichromate ion
3-	HPO ₄ ²⁻	hydrogen phosphate
	PO ₄ ³⁻	phosphate ion

Other Compounds

Formula	Name
H ₂ O	water
NH ₃	ammonia
CH ₄	methane (natural gas)
O ₃	ozone
H ₂ O ₂	hydrogen peroxide
H ₂ , N ₂ , O ₂ , F ₂ , Cl ₂ , Br ₂ , I ₂	diatomic molecules
He, Ne, Ar	Noble gases

Acids

Formula	Name
HNO ₃	nitric acid (strong)
H ₂ SO ₄	sulfuric acid (strong)
HCl	hydrochloric acid (strong)
HBr	hydrobromic acid (strong)
HI	hydroiodic acid (strong)
HClO ₄	perchloric acid (strong)
H ₃ PO ₄	phosphoric acid (weak)
H ₂ CO ₃	carbonic acid (weak)
HC ₂ H ₃ O ₂ (or CH ₃ COOH)	acetic acid (weak)
HF	hydrofluoric acid

Bases

Formula	Name
NaOH	sodium hydroxide (strong)
KOH	potassium hydroxide (strong)
Mg(OH) ₂	magnesium hydroxide (strong)
Ca(OH) ₂	calcium hydroxide (strong)
NH ₃	ammonia (weak)
NR ₃	other amines (weak)

Cst Chemistry Reference Sheet

**American Society for Testing and
Materials**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the text box.

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Book of ASTM Standards, with Related Material American Society for Testing and Materials, 1976 **Annual Book of ASTM Standards** American Society for Testing and Materials, 1978 Index to ASTM standards issued as last part of each vol

EPA-600/2, 1979 Research Reporting Series, 1979 **Tables of Molecular Vibrational Frequencies** Takehiko Shimanouchi, 1972 *Evaluation of Dewatering Devices for Producing High-solids Sludge Cake* Alan F. Cassel, Berinda P. Johnson, 1979 **Next Generation Kinase Inhibitors** Paul Shapiro, 2020-07-14 Protein kinases are fascinating enzymes that maintain the proper function of nearly every task performed by the cells of the human body By extracting a phosphate from the energy molecule ATP and linking it to another protein protein kinases alter the structure and ultimate function of other proteins In this way protein kinases help monitor the extracellular environment and integrate signaling cues that for the most part are beneficial for human health and survival However protein kinases are often dysregulated and responsible for the initiation and progression of many types of cancers inflammatory disorders and other diseases Thus decades of research have revealed much about how protein kinases are regulated and approaches to inhibit these enzymes to treat disease However nearly 30 years since the identification of the first clinically beneficial small molecule protein kinase inhibitor there are only a few examples where these drugs provide sustained and durable patient responses The goal of this book is to provide biomedical scientists graduate and professional degree students insight into different approaches using small molecules to block specific protein kinase functions that promote disease Reader's Index and Guide, 1917 **The Reader's Index & Guide** Croyden Public Libraries, 1930 Reference Catalogue of Current Literature, 1902 **The New Britannica/Webster Dictionary & Reference Guide** Encyclopaedia Britannica, Inc, 1981 *The Library Catalogs of the Hoover Institution on War, Revolution, and Peace, Stanford University* Hoover Institution on War, Revolution, and Peace, 1969 *National Union Catalog*, 1980 Includes entries for maps and atlases *The National Union Catalog, Pre-1956 Imprints*, 1974 **Information Resources in Toxicology** P.J. Bert Hakkinen, Asish Mohapatra, Steven G. G. Gilbert, 2009-08-19 This latest version of Information Resources in Toxicology IRT continues a tradition established in 1982 with the publication of the first edition in presenting an extensive itemization review and commentary on the information infrastructure of the field This book is a unique wide ranging international annotated bibliography and compendium of major resources in toxicology and allied fields such as environmental and occupational health chemical safety and risk assessment Thoroughly updated the current edition analyzes technological changes and is rife with online tools and links to Web sites IRT IV is highly structured providing easy access to its information Among the hot topics covered are Disaster Preparedness and Management Nanotechnology Omics the Precautionary Principle Risk Assessment and Biological Chemical and Radioactive Terrorism and Warfare are among the designated International in scope with contributions from over 30 countries Numerous key references and relevant Web links Concise narratives about toxicologic sub disciplines Valuable

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