

Chemistry Benchmark II Study Guide

1. Write the formula for the following:
 - iron (II) oxide
 - calcium phosphate
 - ammonium phosphate
 - copper (III) bromide
 - nitrogen dioxide
 - titanium (IV) oxide
 - potassium bromide
 - magnesium hydroxide
 - xenon hexafluoride
2. What is a binary compound? What is a ternary compound?
3. Write formulas for:
 - hydroiodic acid
 - carbonic acid
 - hydrosulfuric acid
 - hydrochloric acid
 - hypochlorous acid
 - perchloric acid
 - phosphorous acid
 - chloric acid
 - sulfuric acid
 - chlorous acid
4. Name these:
 - NaCl
 - KI
 - CaBr
 - Rb₂O
 - Sr₂
 - CuCl
 - HgO
 - Fe₂O₃
 - MnO₂
 - PbCl₄
 - SnBr₄
 - BF₃
 - NO
 - N₂O₅
 - CCl₄
 - P₂O₅
 - N₂O₃
5. Find the oxidation number assigned to:
 - S in H₂SO₄
 - P in PCl₅
 - S in SO₃
6. Name these:
 - NO₂
 - N₂O
 - SO₃
7. How many valence electrons do elements have in the following groups?
 - alkali metals
 - alkaline earth metals
 - halogens
 - noble gases
8. Write the complete electron configuration, noble gas configuration, Lewis dot structure, and orbital notation for:
 - Sr
 - Br
 - Na
 - K
 - N
 - Cl
9. In each group, identify the atom with the largest atomic radius.
 - Sn, Xe, Rb, Sr
 - IRn, He, Xe, Kr
 - Pb, Ba, Cs, At
10. In each group, identify the atom with the highest ionization energy.
 - Cs, K, Li
 - Ba, Sr, Ca
 - I, Br, Cl
 - Mg, Si, S
11. Define the following: ionic bond, covalent bond, polar covalent bond, nonpolar covalent bond.
12. Define electronegativity and describe the trend across a period and down a group. What type of element has the highest electronegativity? What type has the lowest electronegativity?
13. Identify the type of bond in each:
 - K-N
 - Cs-O
 - C-N
 - O-F
 - Na-Cl
 - Cl-F
 - Cl-O
 - Ca-Cl
14. Who first arranged the periodic table in order of increasing atomic mass?
15. Draw Lewis structures for the following and predict the molecular shape.
 - H₂S
 - PH₃
 - O₂
 - sulfate ion
 - SO₃
 - BF₃
16. Calculate the wavelength and energy of 96.6MHz radio waves.
17. What is the frequency and energy of electromagnetic radiation having a wavelength of 3.33×10^{-9} m?
18. Predict the ions that the following elements might form:
 - F
 - K
 - I
 - Ge
 - Fe (trick question)
 - P
 - Ag

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