



PERIODIC TABLE OF THE ELEMENTS

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19 K 39.10	20 Ca 40.08	21 Sc 44.96	22 Ti 47.88	23 V 50.94	24 Cr 52.00	25 Mn 54.94	26 Fe 55.85	27 Co 58.93	28 Ni 58.69	29 Cu 63.55	30 Zn 65.38	31 Ga 69.72	32 Ge 72.64	33 As 74.92	34 Se 78.96	35 Br 79.90	36 Kr 83.80	37 Rb 85.47	38 Sr 87.62	39 Y 88.91	40 Zr 91.22	41 Nb 92.91	42 Mo 95.94	43 Tc 98.91	44 Ru 101.07	45 Rh 102.91	46 Pd 106.42	47 Ag 107.87	48 Cd 112.41	49 In 114.82	50 Sn 118.71	51 Sb 121.76	52 Te 127.60	53 I 126.91	54 Xe 131.29	55 Cs 132.91	56 Ba 137.33	57 La 138.91	58 Ce 140.12	59 Pr 140.91	60 Nd 144.24	61 Pm 144.91	62 Sm 150.36	63 Eu 151.96	64 Gd 157.25	65 Tb 158.93	66 Dy 162.50	67 Ho 164.93	68 Er 167.26	69 Tm 168.93	70 Yb 173.05	71 Lu 174.97
87 Fr [223]	88 Ra [226]	89 Ac [227]	90 Th [232]	91 Pa [231]	92 U [238]	93 Np [237]	94 Pu [244]	95 Am [243]	96 Cm [247]	97 Bk [247]	98 Cf [251]	99 Es [252]	100 Fm [257]	101 Md [258]	102 No [259]	103 Lr [262]	104 Rf [261]	105 Db [262]	106 Sg [266]	107 Bh [264]	108 Hs [277]	109 Mt [268]	110 Ds [271]	111 Rg [272]	112 Uub [285]	113 Uut [284]	114 Uuq [289]	115 Uup [288]	116 Uuh [292]	117 Uus [294]	118 Uuo [294]																					

## GENERAL INFORMATION

## A. MATTER

- Chemical changes involve the breaking and/or formation of chemical bonds. Physical changes of substance do not result in the formation of new substances.
- In a reaction, matter is neither created nor destroyed, but only changed from one form to another.
- Mass is a measure of the amount of matter in an object.

## B. ENERGY

- Energy is the ability to do work or transfer heat.
- Energy appears in such forms as light, sound, heat, or mechanical, electrical, or chemical energy.
- Energy can be converted from one form to another, but can be neither created nor destroyed.
- Two general classifications of energy are potential energy and kinetic energy.

C. ONE MOLE OF ANY SUBSTANCE - is that which contains Avogadro's number ( $6.02 \times 10^{23}$ ) of particles. Number of moles = mass/molecular mass.

## ATOMIC STRUCTURE AND THE PERIODIC TABLE

## A. ATOMIC SPECTRA - When an electron moves from the ground state to an excited state, it absorbs energy. When it moves from an excited state to the ground state, it emits energy.

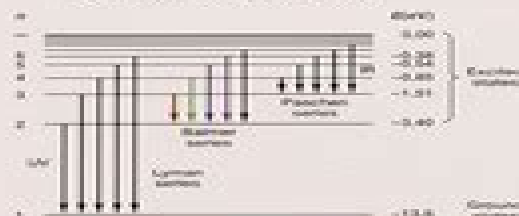
## B. THE BOHR THEORY OF THE HYDROGEN ATOM

- Electrons can exist in only certain stable energy levels.
- When the electronic state of an atom changes, it must absorb or emit exactly that amount of energy equal to the difference between the final and initial states.
- The wavelengths of the lines in the spectrum of hydrogen are given by the Rydberg-Ritz formula:

$$\frac{1}{\lambda} = R \left( \frac{1}{n_1^2} - \frac{1}{n_2^2} \right)$$

where  $R = 109,678 \text{ cm}^{-1}$ ,  $n_1$  and  $n_2$  are quantum numbers for states  $a$  and  $b$ , and  $\lambda$  is the wavelength of light emitted or absorbed.

- Transitions between energy levels are shown by the diagram below.



Energy levels of the hydrogen atom with some of the transitions between them that give rise to the spectral lines indicated.

- Light behaves as if it were composed of quanta of energy known as photons:  $E_{\text{photon}} = h\nu$ , where  $h$  is Planck's constant and  $\nu$  is the frequency of light.

## C. ELECTRIC NATURE OF ATOMS

- The nucleus of an atom is composed of positively charged particles called protons and neutral particles called neutrons.
- The atomic number is the number of protons found in the nucleus.
- The number of protons plus the number of neutrons in the nucleus is the mass number of an isotope of an element.
- In a neutral atom, the number of electrons surrounding the nucleus equals the number of protons in the nucleus.
- The electrons found in the outermost shell are called valence electrons.



## D. THE WAVE MECHANICAL MODEL

- A wave function corresponds to a certain electronic energy and describes a region about the nucleus (orbital) where an electron having that energy may be found.

- The square of the wave function ( $\Psi^2$ ), is called a probability density and equals the probability per unit volume of finding the electron in a given region of space.
- An electron can be described with four quantum numbers that can be used to define allowed energy states.
- The four quantum numbers are principle ( $n$ ), azimuthal ( $l$ ), magnetic ( $m$ ), and spin ( $s$ ).
- Pauli exclusion principle: two electrons within the same atom may not have the same four quantum numbers.

## E. ISOTOPES

- Atoms of the same element having different numbers of neutrons are different isotopes of that element.
- Isotopes of an element have virtually the same chemical properties but different radiochemical properties.
- Stability of the nuclei of an isotope of an element is measured by half-life, which is the time it takes for half the atoms of a sample of an isotope to undergo radioactive decay.
- The atomic mass of an element is calculated by averaging the atomic masses of the isotopes of that element.

## F. PERIODIC TABLE

- The arrangement of the elements in the periodic table reflects the electronic configurations of the atoms of the elements. Each period in the table corresponds to a principle quantum number and a different outer shell that is filling from left to right.
- Group 18 elements are unreactive because their atoms have full, stable outer shells.
- The periodic law states that chemical and physical properties of the elements are periodic functions of their atomic numbers.
- Vertical columns, called groups, consist of elements that have similar chemical properties because they have the same number of valence electrons.
- The most active metals are in the lower left and the most active nonmetals are in the upper right.
- Metallic properties include high electrical conductivity, high melting points, ductility, and malleability.
- Nonmetals are poor electrical conductors and form brittle solids.
- Metalloids have properties intermediate between those of metals and nonmetals.
- Atomic radii generally decrease across a period from left to right and increase down a group.
- Electronegativity generally increases from left to right in a period and decreases down a group.
- Ionization energy decreases going down a group and increases left to right across a period.

## BONDING

## A. BONDS

- An ionic bond occurs when one or more electrons are transferred from the valence shell of one atom to another and there is a coulombic attraction between oppositely charged ions. The atom that loses electrons becomes a positive ion (cation), and the atom that acquires electrons becomes a negatively charged ion (anion).
- The octet rule states that atoms tend to gain or lose electrons until their valence shells are filled.
- In a nonpolar covalent bond, the electrons are shared equally between atoms. In a polar covalent bond, there is unequal sharing.

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### **Table of Contents Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts**

1. Understanding the eBook Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts
  - The Rise of Digital Reading Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts
  - Personalized Recommendations
  - Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts User Reviews and Ratings

- Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts and Bestseller Lists
- 5. Accessing Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts Free and Paid eBooks
  - Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts Public Domain eBooks
  - Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts eBook Subscription Services
  - Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts Budget-Friendly Options
- 6. Navigating Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts Compatibility with Devices
  - Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts
  - Highlighting and Note-Taking Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts
  - Interactive Elements Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts
- 8. Staying Engaged with Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts
- 9. Balancing eBooks and Physical Books Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts
  - Setting Reading Goals Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts
  - Fact-Checking eBook Content of Chemistry Reas Quick Access Reference Chart Quick Access Reference Charts

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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