



CHEMISTRY

PERIODIC TABLE OF THE ELEMENTS

GENERAL INFORMATION

• MATTER

1. Chemical changes involve the breaking and/or formation of chemical bonds. Physical changes of substances do not result in the formation of new substances.
2. In a reaction, matter is neither created nor destroyed, but only changed from one form to another.
3. Mass is a measure of the amount of matter in an object.

REFERENCES

1. Energy is the ability to do work or transfer heat.
2. Energy appears in such forms as light, sound, heat, or mechanical, electrical, or chemical energy.
3. Energy can be converted from one form to another, but can be neither created nor destroyed.
4. Two general classifications of energy are

C. ONE MOLE OF ANY SUBSTANCE – is that which contains Avogadro's number (6.02×10^{23}) of particles. Number of moles = mass/molecular mass.

ATOMIC STRUCTURE AND THE PERIODIC TABLE

ATOMIC SPECTRA

When an electron moves from the ground state to an excited state, it absorbs energy. When it moves from an excited state to the ground state, it emits energy.

B. THE BOHR THEORY OF THE HYDROGEN ATOM

1. Electrons can exist in only certain stable energy levels.
2. When the electronic state of an atom changes, it must absorb or emit exactly that amount of energy equal to the difference between the final and initial states.
3. The wavelengths of the lines in the spectrum of an atom are characteristic of the atom.

the percentage of hydrogen as formaldehyde

where $\beta\hbar = 109,678 \text{ cm}^{-1}$, n_1 and n_2 are quantum numbers for states α and β , and λ is the wavelength of the emitted or absorbed

D. THE WAVE MECHANICAL MODEL

- A wave function corresponds to a certain electronic energy and describes a region about the nucleus (orbital) where an electron having that energy may be found.



更多資訊

2000-2001

A. BONDS

1. An ionic bond occurs when one or more electrons are transferred from the valence shell of one atom to another and there is a coulombic attraction between oppositely charged ions. The atom that loses electrons becomes a positive ion (cation), and the atom that acquires electrons becomes a negatively charged ion (anion).
2. The octet rule states that atoms tend to gain or lose electrons until their valence shells are filled.
3. In a nonpolar covalent bond, the electrons are shared equally between atoms. In a polar covalent bond, there is unequal sharing.

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survivalists with essential fire starting skills for wilderness emergencies This book emphasizes the critical elements of ignition sources tinder selection and stringent safety measures It delves into the science of combustion highlighting how diverse ignition methods from modern tools to primitive techniques can be employed A key insight is the importance of understanding different tinder materials and their properties for effectively catching a spark as well as the absolute necessity of rigorous safety protocols to prevent uncontrolled fires and injuries The book takes a practical hands on approach progressing from basic fire starting principles to advanced techniques Chapters cover identifying and preparing tinder from various environments and tailoring fire types for specific needs like warmth cooking or signaling Grounded in empirical testing field research and referencing data from controlled burns and survival studies Emergency Fire Prep presents information in an actionable format The book integrates knowledge from botany chemistry and environmental science providing a broader understanding of fire's role in different ecosystems This resource stands out by offering clear jargon free directions and checklists that can be directly applied in real world survival situations Focusing exclusively on emergency fire starting in wilderness contexts it serves as a valuable reference work for quick access to critical information making it an essential addition to any emergency preparedness kit

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Carlson,1977 **Information Technology Digest** ,1993 Collier's Encyclopedia Frank Webster Price,1950 Standard Catalog for Public Libraries H.W. Wilson Company,1957 **Standard Catalog for Public Libraries; an Annotated List of 12,300 Titles with a Full Analytical Index; Supplement** H.W. Wilson Company,1954 **Library of Congress Catalog** Library of Congress,1955 A cumulative list of works represented by Library of Congress printed cards Standard Catalog for Public Libraries. Supplement H.W. Wilson Company,1952

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