

**10****CHEMICAL QUANTITIES****Practice Problems**

*On a piece of paper, solve the following problems.*

***YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!***

**SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER**

1. What is the molar mass of sucrose ( $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ )?
2. What is the molar mass of each of the following compounds?
  - a. phosphorus pentachloride ( $\text{PCl}_5$ )
  - b. uranium hexafluoride ( $\text{UF}_6$ )
3. Calculate the molar mass of each of the following ionic compounds:
  - a.  $\text{KMnO}_4$
  - b.  $\text{Ca}_3(\text{PO}_4)_2$
4. How many moles is  $3.52 \times 10^{24}$  molecules of water?
5. How many atoms of zinc are in 0.60 mol of zinc?
6. What is the mass of 1.00 mol of oxygen ( $\text{O}_2$ )?

**SECTION 10.2 MOLE–MASS AND MOLE–VOLUME RELATIONSHIPS**

1. What is the molar mass of each of the following compounds?
  - a.  $\text{C}_6\text{H}_{12}\text{O}_6$
  - b.  $\text{NaHCO}_3$
  - c.  $\text{C}_2\text{H}_2$
  - d.  $\text{KNH}_4\text{SO}_4$
2. Calculate the mass in grams of each of the following:
  - a. 8.0 mol lead oxide ( $\text{PbO}$ )
  - b. 0.75 mol hydrogen sulfide ( $\text{H}_2\text{S}$ )
  - c. 0.00100 mol silicon tetrahydride ( $\text{SiH}_4$ )
  - d.  $1.50 \times 10^{-2}$  mol molecular oxygen ( $\text{O}_2$ )
  - e. 2.30 mol ethylene glycol ( $\text{C}_2\text{H}_6\text{O}_2$ )
3. How many grams are in 1.73 mol of dinitrogen pentoxide ( $\text{N}_2\text{O}_5$ )?
4. How many grams are in 0.658 mol of calcium phosphate [ $\text{Ca}_3(\text{PO}_4)_2$ ]?
5. Calculate the number of moles in each of the following:
  - a. 0.50 g sodium bromide ( $\text{NaBr}$ )
  - b. 13.5 g magnesium nitrate [ $\text{Mg}(\text{NO}_3)_2$ ]
  - c. 1.02 g magnesium chloride ( $\text{MgCl}_2$ )
  - d. 0.00100 g monochloromethane ( $\text{CH}_3\text{Cl}$ )
  - e.  $1.50 \times 10^{-3}$  g propylene glycol [ $\text{C}_3\text{H}_8(\text{OH})_2$ ]
6. A chemist plans to use 435.0 grams of ammonium nitrate ( $\text{NH}_4\text{NO}_3$ ) in a reaction. How many moles of the compound is this?
7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ( $\text{C}_{20}\text{H}_{24}\text{N}_2\text{O}_2$ ). What mass, in grams, should the laboratory technician obtain in order to make the solution?
8. What is the volume at STP of 2.66 mol of methane ( $\text{CH}_4$ ) gas?
9. How many moles is 135 L of ammonia ( $\text{NH}_3$ ) gas at STP?

# Chapter 10 Chemical Quantities Study Guide Answer Key

**Steven S. Zumdahl**



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### **Table of Contents Chapter 10 Chemical Quantities Study Guide Answer Key**

1. Understanding the eBook Chapter 10 Chemical Quantities Study Guide Answer Key
  - The Rise of Digital Reading Chapter 10 Chemical Quantities Study Guide Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 10 Chemical Quantities Study Guide Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 10 Chemical Quantities Study Guide Answer Key
  - User-Friendly Interface



4. Exploring eBook Recommendations from Chapter 10 Chemical Quantities Study Guide Answer Key
  - Personalized Recommendations
  - Chapter 10 Chemical Quantities Study Guide Answer Key User Reviews and Ratings
  - Chapter 10 Chemical Quantities Study Guide Answer Key and Bestseller Lists
5. Accessing Chapter 10 Chemical Quantities Study Guide Answer Key Free and Paid eBooks
  - Chapter 10 Chemical Quantities Study Guide Answer Key Public Domain eBooks
  - Chapter 10 Chemical Quantities Study Guide Answer Key eBook Subscription Services
  - Chapter 10 Chemical Quantities Study Guide Answer Key Budget-Friendly Options
6. Navigating Chapter 10 Chemical Quantities Study Guide Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 10 Chemical Quantities Study Guide Answer Key Compatibility with Devices
  - Chapter 10 Chemical Quantities Study Guide Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 10 Chemical Quantities Study Guide Answer Key
  - Highlighting and Note-Taking Chapter 10 Chemical Quantities Study Guide Answer Key
  - Interactive Elements Chapter 10 Chemical Quantities Study Guide Answer Key
8. Staying Engaged with Chapter 10 Chemical Quantities Study Guide Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 10 Chemical Quantities Study Guide Answer Key
9. Balancing eBooks and Physical Books Chapter 10 Chemical Quantities Study Guide Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 10 Chemical Quantities Study Guide Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 10 Chemical Quantities Study Guide Answer Key
  - Setting Reading Goals Chapter 10 Chemical Quantities Study Guide Answer Key
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chapter 10 Chemical Quantities Study Guide Answer Key
  - Fact-Checking eBook Content of Chapter 10 Chemical Quantities Study Guide Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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