

Chapter Test B

A. Matching

Match each item in Column A with the correct description in Column B. Write the letter of the correct description in the blank.

Column A

- _____ 1. An unknown mass dissolved in a known volume of solvent for solution density is required.
- _____ 2. An unknown volume of gas produced from a known mass of reactant is required.
- _____ 3. The relationship of quantities in chemical equations.
- _____ 4. The ratio of the actual yield to the theoretical yield expressed as a percent.
- _____ 5. The unknown mass of a product or reactant is known and the volume of gas is required.
- _____ 6. An unknown volume of gas produced from a known mass of reactant is required.

Column B

- a. percent yield
- b. limiting reagent
- c. theoretical yield
- d. stoichiometry
- e. percent yield
- f. percent reagent

B. Multiple Choice

Choose the best answer for each question from the list.

- _____ 1. A 10.0-gram sample of a substance is heated and 3.0 grams of gas are evolved. The mass of the residue is _____.
a. 13.0 grams
b. 7.0 grams
c. 10.0 grams
d. 3.0 grams
- _____ 2. Which of the following is a balanced chemical equation for the reaction of hydrogen and oxygen to form water?
a. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
b. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
c. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}_2$
d. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$
- _____ 3. At 100°C, 10.0 g of H_2O are vaporized. The volume of the gas is _____.
a. 10.0 L
b. 10.0 mL
c. 10.0 dm³
d. 10.0 cm³

Chapter 12 Study Guide For Content Mastery

Stoichiometry Key

Johan Roos



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