

## CHAPTER 13

### 13.1 States of matter review

**Gases:** contain particles that are far apart, in random motion, and independent of one another.

- *Little contact*
- *No attractive forces*
- *Random motion*

**Solids:** contain particles very close in space and maintain a rigid shape. Significant attractive forces exist between particles.

- *Close contact*
- *Strong attractive forces*
- *Rigid shape*

Liquids: intermediate between gases and solids.

Contain particles close to one another but have fluidity (can assume the shape of a container).

Significant attractive forces exist between particles in a liquid.

- *Close contact*
- *Some attractive forces*
- *Fluid shape*

## 13.2 Properties of liquids

## Evaporation

### Evaporation or Vaporization:

### *Escape of molecules from the liquid to the gas phase.*

**Kinetic Energy:** Energy associated with motion.

1. Molecules in the liquid state have different kinetic energies (KE's).
2. Those with higher KE's can overcome attractive forces between particles and escape to the gas phase.
3. We can increase the kinetic energy (or movement of the molecules in liquid state by adding heat (boiling).

## Sublimation

Phase change from the solid to gas phase that bypasses the liquid state.

ScHd vapour

### Exemples:

## Fusion

Molecules can go from the liquid to solid phase.

This process is called "fusion".

Ex. Liquid solid

Under certain circumstances, gas can transform directly into a solid. This process is called "deposition".

**Example: Water vapor to ice - Water vapor transforms directly into ice without becoming a liquid, a process that often occurs on windows during the winter months.**

## Condensation

Molecules from the liquid phase can escape to the vapor phase through evaporation.

# Chapter 13 Study Guide Gases Answers

**Arshad Iqbal**

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