

17

THERMOCHEMISTRY

Practice Problems

1. A system absorbs 100 J of heat and does 40 J of work.

SECTION 17.1 THE FLOW OF ENERGY—HEAT AND WORK

1. A system absorbs 100 J of heat and does 40 J of work. Calculate the change in internal energy.

2. A system absorbs 100 J of heat and does 40 J of work. Calculate the change in internal energy.

3. A system absorbs 100 J of heat and does 40 J of work. Calculate the change in internal energy.

4. A system absorbs 100 J of heat and does 40 J of work. Calculate the change in internal energy.

5. A system absorbs 100 J of heat and does 40 J of work. Calculate the change in internal energy.

6. A system absorbs 100 J of heat and does 40 J of work. Calculate the change in internal energy.

SECTION 17.2 MEASURING AND EXPRESSING ENTHALPY CHANGES

1. Calculate the enthalpy change for the reaction of 1 mole of hydrogen gas with 1/2 mole of oxygen gas to form 1 mole of liquid water.

2. Calculate the enthalpy change for the reaction of 1 mole of hydrogen gas with 1/2 mole of oxygen gas to form 1 mole of liquid water.

3. Calculate the enthalpy change for the reaction of 1 mole of hydrogen gas with 1/2 mole of oxygen gas to form 1 mole of liquid water.

4. Calculate the enthalpy change for the reaction of 1 mole of hydrogen gas with 1/2 mole of oxygen gas to form 1 mole of liquid water.

SECTION 17.3 HEAT IN CHANGES OF STATE

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Chris Hren, Peter J. Mikulecky



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