

## 10

## CHEMICAL QUANTITIES

## Practice Problems

On a piece of paper, solve the following problems.

**YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!**

## SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER

1. What is the molar mass of sucrose ( $C_{12}H_{22}O_{11}$ )?
2. What is the molar mass of each of the following compounds?
  - a. phosphorus pentachloride ( $PCl_5$ )
  - b. uranium hexafluoride ( $UF_6$ )
3. Calculate the molar mass of each of the following ionic compounds:
  - a.  $KMnO_4$
  - b.  $Ca_3(PO_4)_2$
4. How many moles is  $3.52 \times 10^{24}$  molecules of water?
5. How many atoms of zinc are in 0.60 mol of zinc?
6. What is the mass of 1.00 mol of oxygen ( $O_2$ )?

## SECTION 10.2 MOLE–MASS AND MOLE–VOLUME RELATIONSHIPS

1. What is the molar mass of each of the following compounds?
  - a.  $C_6H_{12}O_6$
  - b.  $NaHCO_3$
  - c.  $C_2H_{12}$
  - d.  $KNH_4SO_4$
2. Calculate the mass in grams of each of the following:
  - a. 8.0 mol lead oxide ( $PbO$ )
  - b. 0.75 mol hydrogen sulfide ( $H_2S$ )
  - c. 0.00100 mol silicon tetrahydride ( $SiH_4$ )
  - d.  $1.50 \times 10^{-2}$  mol molecular oxygen ( $O_2$ )
  - e. 2.30 mol ethylene glycol ( $C_2H_6O_2$ )
3. How many grams are in 1.73 mol of dinitrogen pentoxide ( $N_2O_5$ )?
4. How many grams are in 0.658 mol of calcium phosphate [ $Ca_3(PO_4)_2$ ]?
5. Calculate the number of moles in each of the following:
  - a. 0.50 g sodium bromide ( $NaBr$ )
  - b. 13.5 g magnesium nitrate [ $Mg(NO_3)_2$ ]
  - c. 1.02 g magnesium chloride ( $MgCl_2$ )
  - d. 0.00100 g monochloromethane ( $CH_3Cl$ )
  - e.  $1.50 \times 10^{-3}$  g propylene glycol [ $C_3H_8(OH)_2$ ]
6. A chemist plans to use 435.0 grams of ammonium nitrate ( $NH_4NO_3$ ) in a reaction. How many moles of the compound is this?
7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ( $C_{20}H_{24}N_2O_2$ ). What mass, in grams, should the laboratory technician obtain in order to make the solution?
8. What is the volume at STP of 2.66 mol of methane ( $CH_4$ ) gas?
9. How many moles is 135 L of ammonia ( $NH_3$ ) gas at STP?

# Chapter 19 Chemistry Study Guide Answers

**Paul A. Bartlett, Judith G. Koch**



## Chapter 19 Chemistry Study Guide Answers:

**Organic Chemistry, Student Study Guide and Solutions Manual** David R. Klein, 2017-01-04 This is the Student Study Guide and Solutions Manual to accompany Organic Chemistry 3e Organic Chemistry 3rd Edition is not merely a compilation of principles but rather it is a disciplined method of thought and analysis Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems Readers must learn to become proficient at approaching new situations methodically based on a repertoire of skills These skills are vital for successful problem solving in organic chemistry Existing textbooks provide extensive coverage of the principles but there is far less emphasis on the skills needed to actually solve problems Organic Chemistry, 5e Student Study Guide and Solutions Manual David R. Klein, Laurie S. Starkey, 2025-03-18 Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems With Organic Chemistry Student Study Guide and Solutions Manual 5th Edition students can learn to become proficient at approaching new situations methodically based on a repertoire of skills These skills are vital for successful problem solving in organic chemistry

**Organic Chemistry Study Guide and Solutions Manual** Marc Loudon, Jim Parise, Lisa Bonner, 2024-08-09 The Study Guide and Solutions Manual offers several study tools for students Study Guide Links are brief supplements to text material and provide hints for study shortcuts approaches to problem solving or additional explanations of topics that typically cause difficulty Further Explorations are short discussions that provide additional depth on content in the text Reaction Reviews present the reactions in the chapter in detailed outline form This summary consolidates the reactions of the chapter in one place and presents not only the reactions themselves but also the essential features of their mechanisms and stereochemistry for cases in which these issues are discussed in the text Solutions are provided for all of the over 1600 problems in the text Many include worked out examples in which not only the answer is given but also the detailed approach to the solution is discussed

**Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition** Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines

*Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl* Paul B. Kelter, 1993

**General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing, 1999 *Organic Chemistry, Study Guide and Solutions Manual* T. W. Graham Solomons, Craig B. Fryhle, 2000 On the cover of this book is a Pacific yew tree found in the ancient forests of the Pacific Northwest The bark of the Pacific yew tree produces Taxol found to be a highly effective drug against ovarian and breast cancer Taxol blocks mitosis during eukaryotic cell division The supply of Taxol from the Pacific yew tree is vanishingly

small however A single 100 year old tree provides only about one dose of the drug roughly 300 mg For this reason as well as the spectacular molecular architecture of Taxol synthetic organic chemists fiercely undertook efforts to synthesize it Five total syntheses of Taxol have thus far been reported Now a combination of isolation of a related metabolite from European yew needles and synthesis of Taxol from that intermediate supply the clinical demand This case clearly demonstrates the importance of synthesis and the use of organic chemistry It s just one of the many examples used in the text that will spark the interest of students and get them involved in the study of organic chemistry

*Fundamentals of Organic Chemistry, Textbook, Study Guide and Solutions Manual* T. W. Graham Solomons, 1996-08-15 A realistic approach to the study of mechanisms The book addresses real functional group chemistry with an emphasis on the biological environmental and medical applications of organic chemistry

**Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed** Robert K. Wismer, 1982

**Study Guide and Solutions Manual for Organic Chemistry** Susan McMurry, 1992 John McMurry s best selling text presents organic chemistry in a new edition that is up to date beautifully written visually striking and pedagogically sound Described by many of its users as an eminently teachable text McMurry sets the standard in the field The writing style has received almost universal acclaim from its users McMurry introduces new concepts only as needed and immediately illustrates them with concrete examples And wherever possible he ties material together with brief reviews overviews and reaction summaries The result is a text that helps students mentally organize the material a text that helps them understand concepts not just memorize facts and a text that helps them make sense of the voluminous amount of material they encounter in the study of organic chemistry McMurry uses a simple but important polar reaction the addition of HBr to an alkene as the lead off reaction to illustrate the general principles of organic reactions Users of former editions found this an excellent choice because of its relative simplicity no prior knowledge of chirality or kinetics is required and its importance as a polar reaction on a common functional group that offers students the key to understanding hundreds of thousands of ionic reactions By selecting this particular model McMurry is able to offer an unusually early presentation of organic reactions

**Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today** Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

**Study Guide with Solutions for Organic Chemistry, Third Edition** Ralph J. Fessenden, Joan S. Fessenden, 1986 Provides a set of additional drill problems chapter by chapter discussions and supplemental instructional material to help students master organic chemistry problem solving techniques

*Organic Chemistry, Study Guide* T. W. Graham Solomons, 1988-03-08 This class tested text reflects the refinements made in previous editions while expanding updating and adding new information on many important

topics Adopting a bio organic emphasis it introduces functional groups early in the text providing an overview of and preparation for subsequent discussions This edition includes increased coverage of Carbon 13 spectra an expanded treatment of conformational effects of molecules with two sections covering basic geometries of molecules and an in depth overview of the subject plus new material on transition metal chemistry and carbohydrate metabolism Worked out examples have been added to chapters on synthesis and end of chapter problems have been expanded by more than 200 The text also includes exceptional full color graphics illustrating conformational effects and general stereochemical properties of organic substances

Chemistry, Student Study Guide John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that's real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You'll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors

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