

Chapter 4 Atomic Structure**Section 4.2 The Structure of an Atom**
(pages 108–112)

This section compares the properties of three subatomic particles. It also discusses atomic numbers, mass numbers, and isotopes.

Reading Strategy (page 108)

Monitoring Your Understanding Before you read, list in the table shown what you know about atoms and what you would like to learn. After you read, list what you have learned. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

What I Know About Atoms	What I Would Like to Learn	What I Have Learned
Most students will know that atoms are the “building blocks” of matter, and some may know that atoms contain subatomic particles.	Based on the title of the section, students may say that they want to learn more about the structure of atoms.	

Properties of Subatomic Particles (pages 108–109)

1. What are three subatomic particles?
a. Protons b. Electrons c. Neutrons
2. Circle the letter that identifies a subatomic particle with a positive charge.
a. nucleus b. proton c. neutron d. electron
3. Why did Chadwick conclude that the particles produced by his experiment were neutral in charge? A charged object did not deflect the paths of the particles.

Comparing Subatomic Particles (pages 109–110)

4. Circle the letters of properties that vary among subatomic particles.
a. color b. mass
 c. charge d. location in the atom
5. Circle the letter of the expression that accurately compares the masses of neutrons and protons.
a. mass of 1 neutron = mass of 1 proton
b. mass of 2000 neutrons = mass of 1 proton
c. mass of 1 electron = mass of 1 proton
d. mass of 1 neutron = mass of 1 electron

Chapter 4 Atomic Structure Answers

Peter P. Berlow

Chapter 4 Atomic Structure Answers:

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Study Guide and Solutions Manual to accompany Basic Concepts of Chemistry, 9th Edition Leo J. Malone,2011-12-07 This is the Study Guide and Solutions Manual to accompany Malone's Basic Concepts of Chemistry Student's Guide to Introduction to Chemical Principles by Edward I. Peters, 2d Ed Peter P. Berlow,1978 MCAT Comprehensive Review ,2004

Instructor's Guide Alan Sherman,Leonard Russikoff,Sharon Sherman,1976 **MATERIALS SCIENCE AND ENGINEERING, SIXTH EDITION** V. RAGHAVAN,2015-05-01 This well established and widely adopted book now in its Sixth Edition provides a thorough analysis of the subject in an easy to read style It analyzes systematically and logically the basic concepts and their applications to enable the students to comprehend the subject with ease The book begins with a clear exposition of the background topics in chemical equilibrium kinetics atomic structure and chemical bonding Then follows a detailed discussion on the structure of solids crystal imperfections phase diagrams solid state diffusion and phase transformations This provides a deep insight into the structural control necessary for optimizing the various properties of materials The mechanical properties covered include elastic anelastic and viscoelastic behaviour plastic deformation creep and fracture phenomena The next four chapters are devoted to a detailed description of electrical conduction superconductivity semiconductors and magnetic and dielectric properties The final chapter on Nanomaterials is an important

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