

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent
insoluble	solid	solvent

After reading (1) _____ solution _____ of solute is dissolved in solvent.

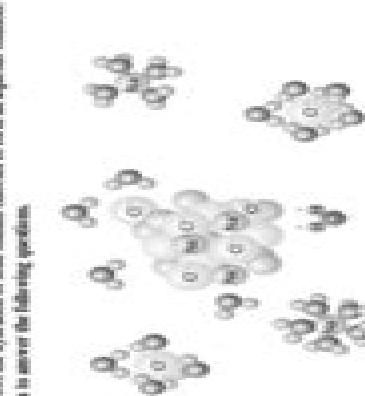
(2) _____ solute _____ is dissolved in the (3) _____ solvent _____.

(4) _____ solubility _____ is the ability of a substance to dissolve in another substance. A substance that does not dissolve is (5) _____ insoluble _____.

(6) _____ liquid _____ solubility of the most common type of solution. If one liquid is soluble in another liquid, such as water and oil, water is the (7) _____ solute _____ and oil is the (8) _____ insoluble _____.

Read about solubility in aqueous solution in your textbook.

Underline the questions to answer the following questions.



1. Indicate in solute is when the solute is in a (8) _____.

Solution is the process of surrounding solute particles with solvent particles to form a solution.

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Section 15.1 content

16. An sodium chloride dissolved in water. What happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

17. Explain the attraction of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atom). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. In your notes, list the uses of solutions may be discussed.

Starting at the top of the solution, breaking the solution into smaller pieces, and heating the solvent.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute/solvent, with type of soln.

True _____ 16. The overall energy change that occurs when a solute dissolves in a solvent is called the heat of solution.

False _____ 17. Solubility is a property of the solute, amount of either the solute or a given amount of solute in a specified temperature and pressure.

False _____ 18. Solubility continues to change as the solute dissolves in a solvent because the heat of solution continues to change.

True _____ 19. In a saturated solution, solute and crystallization are in equilibrium.

True _____ 20. Additional solute can be dissolved in an unsaturated solution, if the solution becomes saturated.

True _____ 21. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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