

Chapter 18 - Reaction Rates and Equilibrium

Collision Theory

- Atoms, ions, and molecules can react to form products when they collide with one another
- Must have sufficient kinetic energy
- Lacking in kinetic energy, particles bounce apart unchanged
- Rate – a measure of speed of any change that occurs within an interval of time



Energy Changes in a Reaction

Activation energy –

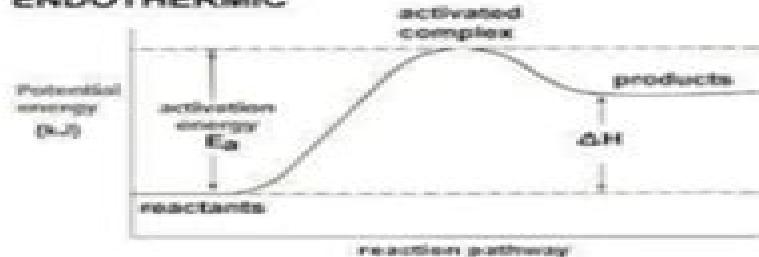
- minimum amount of KE that colliding particles need to react

Activated complex –

- temporary intermediate product
- may reform into the reactants or a new product

Potential Energy Diagram

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Factors Affecting Reaction Rates

1. Effect of Temperature

- Usually increasing temperature speeds up a reaction
- High kinetic energy

2. Effect of Concentration

- Increasing the number of particles into a fixed volume increases the collision frequency

3. Effect of Particle Size

- Increasing surface area increases the collision frequency

4. Effect of Catalysts

- Increases the rate of reaction without being used up itself in the reaction
- Lowers the activation energy



Chapter Assessment Reaction Rates Study Guide

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