

Spring Final Chemistry Study-Guide

General Tips...

- Make sure you look over this WHOLE thing, some answers are given within the review but I will not be giving them to you on the day of the test.
- You will get an unlabeled equation sheet and a periodic table to use.
- **ASSIGNMENT:** you will be assigned to a group to make a review presentation for the class. Each group is responsible for
 1. **A one page handout for the class** that should be printed for the class or posted on-line (e-mail them to me by the due date)
 - Finding/listing all the definitions for their chapter of the review
 - Explaining/listing the equations/variables for that chapter
 - Working out at least one sample problem for each type of problem in your chapter
 - Assigning at least one more practice problem for the class for your chapter
 2. **A presentation to the class:** in the presentation you take your section question by question and provide the information that is necessary for that section. (definition or equation). You do not need to SOLVE any problems but should assign a practice problem for each math related question.

Chapter 10

- 1) Law of Conservation of mass
- 2) Law of conservation of mass, definite proportions (who and what)
- 3) Mole- define multiple ways
- 4) Avogadro's number =
- 5) Molar mass
- 6) Mole \rightarrow grams
- 7) Molecules \rightarrow mol
- 8) Mol \rightarrow grams
- 9) L \rightarrow mol (@STP)
- 10) Mol \rightarrow L
- 11) Standard molar volume
- 12) % percent composition (solve an easy problem)
- 13) Reduce a molecular formula to an empirical formula

Chapter 11

- 14) Balance an equation
- 15) Activity series
- 16) Synthesis, decomposition, single replacement, double replacement, combustion
- 17) Activity series
- 18) Coefficient
- 19) Synthesis, decomposition, single replacement, double replacement, combustion

Chemistry Study Guide Spring 2015

Mary Burnham, Carol Hurd



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