

# Intro. to Chemical Reactions

## Electrolyte Reactions

1. Write down formulas for the following compounds:

a) Barium hydroxide  $\text{Ba}(\text{OH})_2$

b) Sodium phosphate  $\text{Na}_3\text{PO}_4$

c) Magnesium phosphate  $\text{Mg}_3(\text{PO}_4)_2$

d) Calcium iodide  $\text{CaI}_2$

2. Write the formulas for the following acids:

a) HClO<sub>4</sub> Hydrogen perchlorate

b) H<sub>3</sub>PO<sub>4</sub> Hydrogen phosphate

c) H<sub>2</sub>SO<sub>4</sub> Hydrogen sulfate

d) H<sub>2</sub>CO<sub>3</sub> Hydrogen carbonate

## Electrolyte Reactions: Physical Theory, Basic Idea

There is often a chemical reaction between two acids in the following:



There are questions that you can answer about these two aqueous acids:

**Question 1:** What are the products and reactants of this reaction?

There is also an electric relationship between the aqueous acid and aqueous base reaction. This is a short question:

## Corrosive/Neutralizing Reactions

1. Given the following reaction, write the equation that goes along with it:



Corrosive/Neutralizing reactions with either acidic and acidic or basic acidic and basic acidic.

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