

Chemistry Test Cheat Sheet

Combustion Reaction (2 pts)

Example: $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

Balanced: $\text{C}_2\text{H}_6 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

Synthesis Reaction (2 pts)

Example: $\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}$

Balanced: $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$

Limiting/Excess Reactant (4 pts)

Example: 10g H_2 & $80\text{g O}_2 \rightarrow$

H_2 : 4.95 mol, O_2 : 2.50 mol

Need 5 mol H_2 , have 4.95 $\rightarrow \text{H}_2$ is limiting

Ideal Gas Law (4 pts)

$PV = nRT$

Example: $n=1$, $R=0.0821$, $T=273\text{K}$, $V=22.4\text{L} \rightarrow P = 1 \text{ atm}$

% by Mass Solution (4 pts)

Formula: $(\text{solute mass} / \text{solution mass}) \times 100$

Example: 10g NaCl in $90\text{g water} \rightarrow 10\%$

EMR Problem (4 pts)

Formula: $E = hv$, $h = 6.63 \times 10^{-34}$

$v = 5.0 \times 10^{14} \text{ Hz} \rightarrow E = 3.32 \times 10^{-19} \text{ J}$

Noble Gas Configuration (1 pt)

Sulfur (S): $[\text{Ne}] 3s^2 3p^4$

Formula Naming (1 pt ea.)

Na_2SO_4 : Sodium sulfate

CO_2 : Carbon dioxide

FeCl_2 : Iron(II) chloride

Formula Writing (1 pt ea.)

Calcium fluoride $\rightarrow \text{CaF}_2$

Ammonium phosphate $\rightarrow (\text{NH}_4)_3\text{PO}_4$

Dinitrogen pentoxide $\rightarrow \text{N}_2\text{O}_5$

VSEPR Shape (5 pts)

Example: $\text{CH}_4 \rightarrow \text{Tetrahedral}$, 109.5°

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DP Hallahan



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