

# CHAPTER 8 STUDY GUIDE

## Covalent Bonding

### Section 8.1 The Covalent Bond

*In your textbook, read about the nature of covalent bonds.*

Use each of the terms below just once to complete the passage.

covalent bond	molecule	sigma bond	exothermic	pi bond
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When sharing of electrons occurs, the attachment between atoms that results is called a(n) **(1)** \_\_\_\_\_. When such an attachment is formed, bond dissociation energy is released, and the process is **(2)** \_\_\_\_\_. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **(3)** \_\_\_\_\_. If the electrons shared are centered between the two atoms, the attachment is called a(n) **(4)** \_\_\_\_\_. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **(5)** \_\_\_\_\_.

*In your textbook, read about single and multiple bonds and bond strength.*

Circle the letter of the choice that best completes the statement or answers the question.

- In what form do elements such as hydrogen, nitrogen, and oxygen normally occur?
  - as single atoms
  - as molecules containing two atoms
  - as molecules containing three atoms
  - as molecules containing four atoms
- How many electrons are shared in a double covalent bond?
  - none
  - one
  - two
  - four
- Bond length is the distance between
  - two molecules of the same substance.
  - the electrons in two attached atoms.
  - the nuclei of two attached atoms.
  - the orbitals of two attached atoms.
- Which of the following relationships relating to bond length is generally correct?
  - the shorter the bond, the stronger the bond
  - the shorter the bond, the weaker the bond
  - the shorter the bond, the fewer the electrons in it
  - the shorter the bond, the lower the bond dissociation energy

# Covalent Bonds Study Guide

**David R. Klein, Laurie S. Starkey**

A red circular graphic with a gradient, appearing as a partial circle on the right side of the slide, overlapping the light blue bar.

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